

(12) INTERNATIONAL APPLICATION PUBLISHED UNDER THE PATENT COOPERATION TREATY (PCT)

(19) World Intellectual Property Organization  
International Bureau



(43) International Publication Date  
31 October 2002 (31.10.2002)

(10) International Publication Number  
PCT WO 02/086223 A1

(51) International Patent Classification<sup>7</sup>: D06L 1/02,  
C11D 7/50, B08B 7/00

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(21) International Application Number: PCT/US02/12304

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(22) International Filing Date: 18 April 2002 (18.04.2002)

(25) Filing Language: English

(26) Publication Language: English

(30) Priority Data:  
09/837,849 18 April 2001 (18.04.2001) US

(81) Designated States (*national*): AE, AG, AL, AM, AT, AU,  
AZ, BA, BB, BG, BR, BY, BZ, CA, CH, CN, CO, CR, CU,  
CZ, DE, DK, DM, DZ, EC, EE, ES, FI, GB, GD, GE, GH,  
GM, HR, HU, ID, IL, IN, IS, JP, KE, KG, KP, KR, KZ, LC,  
LK, LR, LS, LT, LU, LV, MA, MD, MG, MK, MN, MW,  
MX, MZ, NO, NZ, OM, PH, PL, PT, RO, RU, SD, SE, SG,  
SI, SK, SL, TJ, TM, TN, TR, TT, TZ, UA, UG, UZ, VN,  
YU, ZA, ZM, ZW.

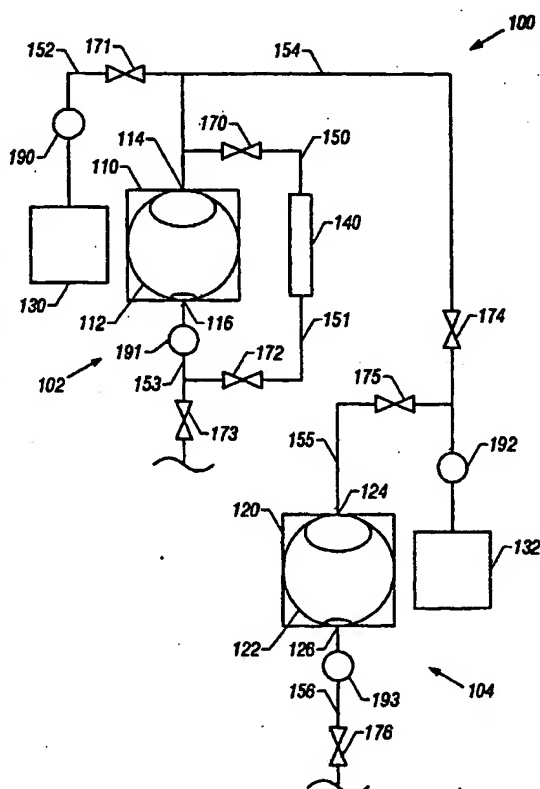
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(84) Designated States (*regional*): ARIPO patent (GH, GM,  
KE, LS, MW, MZ, SD, SL, SZ, TZ, UG, ZM, ZW),  
Eurasian patent (AM, AZ, BY, KG, KZ, MD, RU, TJ, TM),  
European patent (AT, BE, CH, CY, DE, DK, ES, FI, FR,

[Continued on next page]

(54) Title: CLEANING SYSTEM UTILIZING AN ORGANIC CLEANING SOLVENT AND A PRESSURIZED FLUID SOL-  
VENT



(57) Abstract: A cleaning system that utilizes an organic cleaning solvent and pressurized fluid solvent is disclosed. The system has no conventional evaporative hot air drying cycle. Instead, the system utilizes the solubility of the organic solvent in pressurized fluid solvent as well as the physical properties of pressurized fluid solvent. After an organic solvent cleaning cycle, the solvent is extracted from the textiles at high speed in a rotating drum (112, 122) in the same way conventional solvents are extracted from textiles in conventional evaporative hot air dry cleaning machines. Instead of proceeding to a conventional drying cycle, the extracted textiles are then immersed in pressurized fluid solvent to extract the residual organic solvent from the textiles. This is possible because the organic solvent is soluble in pressurized fluid solvent. After the textiles are immersed in pressurized fluid solvent, pressurized fluid solvent is pumped from the drum (112, 122). Finally, the drum is de-pressurized to atmospheric pressure to evaporate any remaining pressurized fluid solvent, yielding clean, solvent free textiles. The organic solvent is preferably selected from terpenes, halohydrocarbons, certain glycol ethers, polyols, ethers, esters of glycol ethers, esters of fatty acids and other long chain carboxylic acids, fatty alcohols and other long-chain alcohols, short-chain alcohols, polar aprotic solvents, siloxanes, hydrofluoroethers, dibasic esters, and aliphatic hydrocarbons solvents or similar solvents or mixtures of such solvents and the pressurized fluid solvent is preferably densified carbon dioxide.

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GB, GR, IE, IT, LU, MC, NL, PT, SE, TR), OAPI patent  
(BF, BJ, CF, CG, CI, CM, GA, GN, GQ, GW, ML, MR,  
NE, SN, TD, TG).

— before the expiration of the time limit for amending the  
claims and to be republished in the event of receipt of  
amendments

**Published:**

— with international search report

For two-letter codes and other abbreviations, refer to the "Guid-  
ance Notes on Codes and Abbreviations" appearing at the begin-  
ning of each regular issue of the PCT Gazette.

## CLEANING SYSTEM UTILIZING AN ORGANIC CLEANING SOLVENT AND A PRESSURIZED FLUID SOLVENT

### BACKGROUND

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#### Field of the Invention

The present invention relates generally to cleaning systems, and more specifically to substrate cleaning systems, such as textile cleaning systems, utilizing an organic cleaning solvent and a pressurized fluid solvent.

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#### Related Art

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A variety of methods and systems are known for cleaning substrates such as textiles, as well as other flexible, precision, delicate, or porous structures that are sensitive to soluble and insoluble contaminants. These known methods and systems typically use water, perchloroethylene, petroleum, and other solvents that are liquid at or substantially near atmospheric pressure and room temperature for cleaning the substrate.

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Such conventional methods and systems generally have been considered satisfactory for their intended purpose. Recently, however, the desirability of employing these conventional methods and systems has been questioned due to environmental, hygienic, occupational hazard, and waste disposal concerns, among other things. For example, perchloroethylene frequently is used as a solvent to clean delicate substrates, such as textiles, in a process referred to as "dry cleaning." Some locales require that the use and disposal of this solvent be regulated by environmental agencies, even when only trace amounts of this solvent are to be introduced into waste streams.

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Furthermore, there are significant regulatory burdens placed on solvents such as perchloroethylene by agencies such as the EPA, OSHA and DOT. Such regulation results in increased costs to the user, which, in turn, are passed to the ultimate consumer. For example, filters that have been used in conventional perchloroethylene dry cleaning systems must be disposed of in accordance with hazardous waste or other environmental regulations. Certain other solvents used in dry cleaning, such as hydrocarbon solvents, are extremely flammable, resulting in greater occupational hazards to the user and increased costs to control their use.

In addition, textiles that have been cleaned using conventional cleaning methods are typically dried by circulating hot air through the textiles as they are tumbled in a drum. The solvent must have a relatively high vapor pressure and low boiling point to be used effectively in a system utilizing hot air drying. The heat used in drying may permanently set some stains in the textiles. Furthermore, the drying cycle adds significant time to the overall processing time. During the conventional drying process, moisture adsorbed on the textile fibers is often removed in addition to the solvent. This often results in the development of undesirable static electricity and shrinkage in the garments. Also, the textiles are subject to greater wear due to the need to tumble the textiles in hot air for a relatively long time. Conventional drying methods are inefficient and often leave excess residual solvent in the textiles, particularly in heavy textiles, components constructed of multiple fabric layers, and structural components of garments such as shoulder pads. This may result in unpleasant odors and, in extreme cases, may cause irritation to the skin of the wearer. In addition to being time consuming and of limited efficiency, conventional drying results in significant loss of cleaning solvent in the form of fugitive solvent vapor. The heating required to evaporate combustible solvents in a conventional drying process increases the risk of fire and/or explosions. In many cases, heating the solvent will necessitate explosion-proof components and other expensive safety devices to minimize the risk of fire and explosions. Finally, conventional hot air drying is an energy intensive process that results in relatively high utility costs and accelerated equipment wear.

Traditional cleaning systems may utilize distillation in conjunction with filtration and adsorption to remove soils dissolved and suspended in the cleaning solvent. The filters and adsorptive materials become saturated with solvent, therefore, disposal of some filter waste is regulated by state or federal laws. Solvent evaporation especially during the drying cycle is one of the main sources of solvent loss in conventional systems. Reducing solvent loss improves the environmental and economic aspects of cleaning substrates using cleaning solvents. It is therefore advantageous to provide a method and system for cleaning substrates that utilizes a solvent having less adverse attributes than those solvents currently used and reduces solvent losses.

As an alternative to conventional cleaning solvents, pressurized fluid solvents or densified fluid solvents have been used for cleaning various substrates, wherein

densified fluids are widely understood to encompass gases that are pressurized to either subcritical or supercritical conditions so as to achieve a liquid or a supercritical fluid having a density approaching that of a liquid. In particular, some patents have disclosed the use of a solvent such as carbon dioxide that is maintained in a liquid state or either a subcritical or supercritical condition for cleaning such substrates as textiles, as well as other flexible, precision, delicate, or porous structures that are sensitive to soluble and insoluble contaminants.

For example, U.S. Patent No. 5,279,615 discloses a process for cleaning textiles using densified carbon dioxide in combination with a non-polar cleaning adjunct. The preferred adjuncts are paraffin oils such as mineral oil or petrolatum. These substances are a mixture of alkanes including a portion of which are C<sub>16</sub> or higher hydrocarbons. The process uses a heterogeneous cleaning system formed by the combination of the adjunct which is applied to the textile prior to or substantially at the same time as the application of the densified fluid. According to the data disclosed in Patent No. 5,279,615, the cleaning adjunct is not as effective at removing soil from fabric as conventional cleaning solvents or as the solvents described for use in the present invention as disclosed below.

U.S. Patent No. 5,316,591 discloses a process for cleaning substrates using liquid carbon dioxide or other liquefied gases below their critical temperature. The focus of this patent is on the use of any one of a number of means to effect cavitation to enhance the cleaning performance of the liquid carbon dioxide. In all of the disclosed embodiments, densified carbon dioxide is the cleaning medium. This patent does not describe the use of a solvent other than the liquefied gas for cleaning substrates. While the combination of ultrasonic cavitation and liquid carbon dioxide may be well suited to processing complex hardware and substrates containing extremely hazardous contaminants, this process is too costly for the regular cleaning of textile substrates. Furthermore, the use of ultrasonic cavitation is less effective for removing contaminants from textiles than it is for removing contaminants from hard surfaces.

U.S. Patent No. 5,377,705, issued to Smith et al., discloses a system designed to clean parts utilizing supercritical carbon dioxide and an environmentally friendly co-solvent. Parts to be cleaned are placed in a cleaning vessel along with the co-solvent. After adding super critical carbon dioxide, mechanical agitation is applied via sonication or brushing. Loosened contaminants are then flushed from

the cleaning vessel using additional carbon dioxide. Use of this system in the cleaning of textiles is neither suggested nor disclosed. Furthermore, use of this system for the cleaning of textiles would result in redeposition of loosened soil and damage to some fabrics.

5 U.S. Patent No. 5,417,768, issued to Smith et al., discloses a process for precision cleaning of a work piece using a multi-solvent system in which one of the solvents is liquid or supercritical carbon dioxide. The process results in minimal mixing of the solvents and incorporates ultrasonic cavitation in such a way as to prevent the ultrasonic transducers from coming in contact with cleaning solvents that  
10 could degrade the piezoelectric transducers. Use of this system in the cleaning of textiles is neither suggested nor disclosed. In fact, its use in cleaning textiles would result in redeposition of loosened soil and damage to some fabrics.

U.S. Patent No. 5,888,250 discloses the use of a binary azeotrope comprised of propylene glycol tertiary butyl ether and water as an environmentally attractive  
15 replacement for perchlorethylene in dry cleaning and degreasing processes. While the use of propylene glycol tertiary butyl ether is attractive from an environmental regulatory point of view, its use as disclosed in this invention is in a conventional dry cleaning process using conventional dry cleaning equipment and a conventional evaporative hot air drying cycle. As a result, it has many of the same disadvantages  
20 as conventional dry cleaning processes described above.

U.S. Patent No. 6,200,352 discloses a process for cleaning substrates in a cleaning mixture comprising carbon dioxide, water, surfactant, and organic co-solvent. This process uses carbon dioxide as the primary cleaning media with the other components included to enhance the overall cleaning effectiveness of the  
25 process. There is no suggestion of a separate, low pressure cleaning step followed by the use of densified fluid to remove the cleaning solvent. As a result, this process has many of the same cost and cleaning performance disadvantages of other liquid carbon dioxide cleaning processes. Additional patents have been issued to the assignee of U.S. Patent No. 6,200,352 covering related subject matter. All of these  
30 patents disclose processes in which liquid carbon dioxide is the cleaning solvent. Consequently, these processes have the same cost and cleaning performance disadvantages.

Several of the pressurized fluid solvent cleaning methods described in the above patents may lead to recontamination of the substrate and degradation of

efficiency because the contaminated solvent is not continuously purified or removed from the system. Furthermore, pressurized fluid solvent alone is not as effective at removing some types of soil as are conventional cleaning solvents. Consequently, pressurized fluid solvent cleaning methods require individual treatment of stains and heavily soiled areas of textiles, which is a labor-intensive process. Furthermore, systems that utilize pressurized fluid solvents for cleaning are more expensive and complex to manufacture and maintain than conventional cleaning systems. Finally, few if any conventional surfactants can be used effectively in pressurized fluid solvents. The surfactants and additives that can be used in pressurized fluid solvent cleaning systems are much more expensive than those used in conventional cleaning systems.

There thus remains a need for an efficient and economic method and system for cleaning substrates that incorporates the benefits of prior systems, and minimizes the difficulties encountered with each. There also remains a need for a method and system in which the hot air drying time is eliminated, or at least reduced, thereby reducing the wear on the substrate and preventing stains from being permanently set on the substrate.

#### SUMMARY

In the present invention, certain types of organic solvents, such as terpenes, halohydrocarbons, certain glycol ethers, polyols, ethers, esters of glycol ethers, esters of fatty acids and other long chain carboxylic acids, fatty alcohols and other long-chain alcohols, short-chain alcohols, polar aprotic solvents, siloxanes, hydrofluoroethers, dibasic esters, and aliphatic hydrocarbons solvents or similar solvents or mixtures of such solvents are used in cleaning substrates. Any type of organic solvent that falls within the range of properties disclosed hereinafter may be used to clean substrates. However, unlike conventional cleaning systems, in the present invention, a conventional drying cycle is not performed. Instead, the system utilizes the solubility of the organic solvent in pressurized fluid solvents, as well as the physical properties of pressurized fluid solvents, to dry the substrate being cleaned.

As used herein, the term "pressurized fluid solvent" refers to both pressurized liquid solvents and densified fluid solvents. The term "pressurized liquid solvent" as used herein refers to solvents that are liquid at between approximately 600 and 1050

pounds per square inch and between approximately 5 and 30 degrees Celsius, but are gas at atmospheric pressure and room temperature. The term "densified fluid solvent" as used herein refers to a gas or gas mixture that is compressed to either subcritical or supercritical conditions so as to achieve either a liquid or a supercritical fluid having density approaching that of a liquid. Preferably, the pressurized fluid solvent used in the present invention is an inorganic substance such as carbon dioxide, xenon, nitrous oxide, or sulfur hexafluoride. Most preferably, the pressurized fluid solvent is densified carbon dioxide.

The substrates are cleaned in a perforated drum within a vessel in a cleaning cycle using an organic solvent. A perforated drum is preferred to allow for free interchange of solvent between the drum and vessel as well as to transport soil from the substrates to the filter. After substrates have been cleaned in the perforated drum, the organic solvent is extracted from the substrates by rotating the cleaning drum at high speed within the cleaning vessel in the same way conventional solvents are extracted from substrates in conventional cleaning machines. However, instead of proceeding to a conventional evaporative hot air drying cycle, the substrates are immersed in pressurized fluid solvent to extract the residual organic solvent from the substrates. This is possible because the organic solvent is soluble in the pressurized fluid solvent. After the substrates are immersed in pressurized fluid solvent, the pressurized fluid solvent is transferred from the drum. Finally, the vessel is de-pressurized to atmospheric pressure to evaporate any remaining pressurized fluid solvent, yielding clean, solvent-free substrates.

The solvents used in the present invention tend to be soluble in pressurized fluid solvents such as supercritical or subcritical carbon dioxide so that a conventional hot air drying cycle is not necessary. The types of solvents used in conventional cleaning systems must have reasonably high vapor pressures and low boiling points because they must be removed from the substrates by evaporation in a stream of hot air. However, solvents that have a high vapor pressure and a low boiling point generally also have a low flash point. From a safety standpoint, organic solvents used in cleaning substrates should have a flash point that is as high as possible, or preferably, it should have no flash point. By eliminating the conventional hot air evaporative drying process, a wide range of solvents can be used in the present invention that have much lower evaporation rates, higher boiling points and higher flash points than those used in conventional cleaning systems. For situations



where the desired solvent has a relatively low flash point, the elimination of the hot air evaporative drying cycle significantly increases the level of safety with respect to fire and explosions.

Thus, the cleaning system described herein utilizes solvents that are less regulated and less combustible, and that efficiently remove different soil types typically deposited on textiles through normal use. The cleaning system reduces solvent consumption and waste generation as compared to conventional dry cleaning systems. Machine and operating costs are reduced as compared to currently used pressurized fluid solvent systems, and conventional additives may be used in the cleaning system.

Furthermore, one of the main sources of solvent loss from conventional dry cleaning systems, which occurs in the evaporative hot air drying step, is substantially reduced or eliminated altogether. Because the conventional evaporative hot air drying process is eliminated, there are no heat set stains on the substrates, risk of fire and/or explosion is reduced, the cleaning cycle time is reduced, and residual solvent in the substrates is substantially reduced or eliminated. Substrates are also subject to less wear, less static electricity build-up and less shrinkage because there is no need to tumble the substrates in a stream of hot air to dry them.

While systems according to the present invention utilizing pressurized fluid solvent to remove organic solvent can be constructed as wholly new systems, existing conventional solvent systems can also be converted to utilize the present invention. An existing conventional solvent system can be used to clean substrates with organic solvent, and an additional pressurized chamber for drying substrates with pressurized fluid solvent can be added to the existing system.

Therefore, according to the present invention, textiles to be cleaned are placed in a cleaning drum within a cleaning vessel, adding an organic solvent to the cleaning vessel, cleaning the textiles with the organic solvent, removing a portion of the organic solvent from the cleaning vessel, rotating the cleaning drum to extract a portion of the organic solvent from the textiles, placing the textiles into a drying drum within a pressurizable drying vessel, adding a pressurized fluid solvent to the drying vessel, removing a portion of the pressurized fluid solvent from the drying vessel, rotating the drying drum to extract a portion of the pressurized fluid solvent from the textiles, depressurizing the drying vessel to remove the remainder of the pressurized

fluid solvent by evaporation, and removing the textiles from the depressurized vessel.

These and other features and advantages of the invention will be apparent upon consideration of the following detailed description of the presently preferred embodiment of the invention, taken in conjunction with the claims and appended drawings, as well as will be learned by practice of the invention.

### BRIEF DESCRIPTION OF THE DRAWINGS

FIG. 1 is a block diagram of a cleaning system utilizing separate vessels for cleaning and drying.

FIG. 2 is a block diagram of a cleaning system utilizing a single vessel for cleaning and drying.

### DETAILED DESCRIPTION

Reference will now be made in detail to embodiments of the invention, examples of which are illustrated in the accompanying drawings. The steps of each method for cleaning and drying a substrate will be described in conjunction with the detailed description of the system.

The methods and systems presented herein may be used for cleaning a variety of substrates. The present invention is particularly suited for cleaning substrates such as textiles, as well as other flexible, precision, delicate, or porous structures that are sensitive to soluble and insoluble contaminants. The term "textile" is inclusive of, but not limited to, woven or non-woven materials, as well as articles made therefrom. Textiles include, but are not limited to, fabrics, articles of clothing, protective covers, carpets, upholstery, furniture and window treatments. For purposes of explanation and illustration, and not limitation, exemplary embodiments of a system for cleaning textiles in accordance with the invention are shown in FIGS. 1 and 2.

As noted above, the pressurized fluid solvent used in the present invention is either a pressurized liquid solvent or a densified fluid solvent. Although a variety of solvents may be used, it is preferred that an inorganic substance such as carbon dioxide, xenon, nitrous oxide, or sulfur hexafluoride, be used as the pressurized fluid solvent. For cost and environmental reasons, liquid, supercritical, or subcritical carbon dioxide is the preferred pressurized fluid solvent.

Furthermore, to maintain the pressurized fluid solvent in the appropriate fluid state, the internal temperature and pressure of the system must be appropriately controlled relative to the critical temperature and pressure of the pressurized fluid solvent. For example, the critical temperature and pressure of carbon dioxide is approximately 31 degrees Celsius and approximately 73 atmospheres, respectively. The temperature may be established and regulated in a conventional manner, such as by using a heat exchanger in combination with a thermocouple or similar regulator to control temperature. Likewise, pressurization of the system may be performed using a pressure regulator and a pump and/or compressor in combination with a pressure gauge. These components are conventional and are not shown in FIGS. 1 and 2 as placement and operation of these components are known in the art.

The system temperature and pressure may be monitored and controlled either manually, or by a conventional automated controller (which may include, for example, an appropriately programmed computer or appropriately constructed microchip) that receives signals from the thermocouple and pressure gauge, and then sends corresponding signals to the heat exchanger and pump and/or compressor, respectively. Unless otherwise noted, the temperature and pressure is appropriately maintained throughout the system during operation. As such, elements contained within the system are constructed of sufficient size and material to withstand the temperature, pressure, and flow parameters required for operation, and may be selected from, or designed using, any of a variety of presently available high pressure hardware.

In the present invention, the preferred organic solvent should have a flash point of greater than 100 F to allow for increased safety and less governmental regulation, have a low evaporation rate to minimize fugitive emissions, be able to remove soils consisting of insoluble particulate soils and solvent soluble oils and greases, and prevent or reduce redeposition of soil onto the textiles being cleaned. Preferably, the organic solvents suitable for use in the present invention include any of the following alone or in combination:

1. *Cyclic terpenes*, specifically,  $\alpha$ -terpene isomers, pine oil,  $\alpha$ -pinene isomers, and d-limonene. Additionally, any cyclic terpene exhibiting the following physical characteristics is suitable for use in the present invention; (1) soluble

in carbon dioxide at a pressure of between 600 and about 1050 pounds per square inch and at a temperature of between 5 and about 30 degrees Celsius; (2) specific gravity of greater than about 0.800 (the higher the specific gravity the better the organic solvent); (3) Hansen solubility parameters of about 13.0 – 17.5 (MPa)<sup>1/2</sup> for dispersion, about 0.5 – 9.0 (MPa)<sup>1/2</sup> for polar, and about 0.0 – 10.5 (MPa)<sup>1/2</sup> for hydrogen bonding.

2. *Halocarbons*, specifically, chlorinated, fluorinated and brominated hydrocarbons exhibiting the following physical characteristics; (1) soluble in carbon dioxide at a pressure of between 600 and about 1050 pounds per square inch and at a temperature of between 5 and about 30 degrees Celsius; (2) specific gravity of greater than about 1.100 (the higher the specific gravity the better the organic solvent); (3) Hansen solubility parameters of about 10.0 – 17.0 (MPa)<sup>1/2</sup> for dispersion, about 0.0 – 7.0 (MPa)<sup>1/2</sup> for polar, and about 0.0 – 5.0 (MPa)<sup>1/2</sup> for hydrogen bonding.

3. *Glycol ethers*, specifically, mono-, di-, triethylene and mono-, di- and tripropylene glycol ethers exhibiting the following physical characteristics; (1) soluble in carbon dioxide at a pressure of between 600 and about 1050 pounds per square inch and at a temperature of between 5 and about 30 degrees Celsius; (2) specific gravity of greater than about 0.800 (the higher the specific gravity the better the organic solvent); (3) Hansen solubility parameters of about 13.0 – 19.5 (MPa)<sup>1/2</sup> for dispersion, about 3.0 – 7.5 (MPa)<sup>1/2</sup> for polar, and about 8.0 – 17.0 (MPa)<sup>1/2</sup> for hydrogen bonding.

4. *Polyols*, specifically, glycols and other organic compounds containing two or more hydroxyl radicals and exhibiting the following physical characteristics; (1) soluble in carbon dioxide at a pressure of between 600 and about 1050 pounds per square inch and at a temperature of between 5 and about 30 degrees Celsius; (2) specific gravity of greater than about 0.920 (the higher the specific gravity the better the organic solvent); (3) Hansen solubility parameters of about 14.0 – 18.2 (MPa)<sup>1/2</sup> for dispersion, about 4.5 – 20.5 (MPa)<sup>1/2</sup> for polar, and about 15.0 – 30.0 (MPa)<sup>1/2</sup> for hydrogen bonding.

5. *Ethers*, specifically, ethers containing no free hydroxyl radicals and exhibiting the following physical characteristics; (1) soluble in carbon dioxide at a pressure of between 600 and about 1050 pounds per square inch and at a temperature of between 5 and about 30 degrees Celsius; (2) specific gravity of greater than about 0.800 (the higher the specific gravity the better the organic solvent); (3) Hansen solubility parameters of about 14.5 – 20.0 (MPa)<sup>1/2</sup> for dispersion, about 1.5 – 6.5 (MPa)<sup>1/2</sup> for polar, and about 5.0 – 10.0 (MPa)<sup>1/2</sup> for hydrogen bonding.
6. *Esters of glycol ethers*, specifically, esters of glycol ethers exhibiting the following physical characteristics; (1) soluble in carbon dioxide at a pressure of between 600 and about 1050 pounds per square inch and at a temperature of between 5 and about 30 degrees Celsius; (2) specific gravity of greater than about 0.800 (the higher the specific gravity the better the organic solvent); (3) Hansen solubility parameters of about 15.0 – 20.0 (MPa)<sup>1/2</sup> for dispersion, about 3.0 – 10.0 (MPa)<sup>1/2</sup> for polar, and about 8.0 – 16.0 (MPa)<sup>1/2</sup> for hydrogen bonding.
7. *Esters of monobasic carboxylic acids* exhibiting the following physical characteristics; (1) soluble in carbon dioxide at a pressure of between 600 and about 1050 pounds per square inch and at a temperature of between 5 and about 30 degrees Celsius; (2) specific gravity of greater than about 0.800 (the higher the specific gravity the better the organic solvent); (3) Hansen solubility parameters of about 13.0 – 17.0 (MPa)<sup>1/2</sup> for dispersion, about 2.0 – 7.5 (MPa)<sup>1/2</sup> for polar, and about 1.5 – 6.5 (MPa)<sup>1/2</sup> for hydrogen bonding.
8. *Fatty alcohols*, specifically alcohols in which the carbon chain adjacent to the hydroxyl group contains five carbon atoms or more and exhibiting the following physical characteristics; (1) soluble in carbon dioxide at a pressure of between 600 and about 1050 pounds per square inch and at a temperature of between 5 and about 30 degrees Celsius; (2) specific gravity of greater than about 0.800 (the higher the specific gravity the better the organic solvent); (3) Hansen solubility parameters of about 13.3 – 18.4 (MPa)<sup>1/2</sup> for

dispersion, about  $3.1 - 18.8 \text{ (MPa)}^{1/2}$  for polar, and about  $8.4 - 22.3 \text{ (MPa)}^{1/2}$  for hydrogen bonding.

- 5 9. *Short chain alcohols* in which the carbon chain adjacent to the hydroxyl group contains four or fewer carbon atoms and exhibiting the following physical characteristics; (1) soluble in carbon dioxide at a pressure of between 600 and about 1050 pounds per square inch and at a temperature of between 5 and about 30 degrees Celsius; (2) specific gravity of greater than about 0.800 (the higher the specific gravity the better the organic solvent); (3) Hansen solubility parameters of about  $13.5 - 18.0 \text{ (MPa)}^{1/2}$  for dispersion, about  $3.0 - 9.0 \text{ (MPa)}^{1/2}$  for polar, and about  $9.0 - 16.5 \text{ (MPa)}^{1/2}$  for hydrogen bonding.
- 10 10. *Siloxanes* exhibiting the following physical characteristics; (1) soluble in carbon dioxide at a pressure of between 600 and about 1050 pounds per square inch and at a temperature of between 5 and about 30 degrees Celsius; (2) specific gravity of greater than about 0.900 (the higher the specific gravity the better the organic solvent); (3) Hansen solubility parameters of about  $14.0 - 18.0 \text{ (MPa)}^{1/2}$  for dispersion, about  $0.0 - 4.5 \text{ (MPa)}^{1/2}$  for polar, and about  $0.0 - 4.5 \text{ (MPa)}^{1/2}$  for hydrogen bonding.
- 15 20 11. *Hydrofluoroethers* exhibiting the following physical characteristics; (1) soluble in carbon dioxide at a pressure of between 600 and about 1050 pounds per square inch and at a temperature of between 5 and 30 degrees Celsius; (2) specific gravity of greater than about 1.50; (3) total Hansen solubility parameters of about  $12.0 - 18.0 \text{ (MPa)}^{1/2}$  for dispersion, about  $4.0 - 10.0 \text{ (MPa)}^{1/2}$  for polar, and about  $1.5 - 9.0 \text{ (MPa)}^{1/2}$  for hydrogen bonding.
- 25 30 12. *Aliphatic hydrocarbons* exhibiting the following physical characteristics; (1) soluble in carbon dioxide at a pressure of between 600 and about 1050 pounds per square inch and at a temperature of between 5 and about 30 degrees Celsius; (2) specific gravity of greater than about 0.700 (the higher the specific gravity the better the organic solvent); (3) Hansen solubility parameters of about  $14.0 - 17.0 \text{ (MPa)}^{1/2}$  for dispersion, about  $0.0 - 2.0 \text{ (MPa)}^{1/2}$  for polar, and about  $0.0 - 2.0 \text{ (MPa)}^{1/2}$  for hydrogen bonding.

- 5 13. *Esters of dibasic carboxylic acids* exhibiting the following physical characteristics; (1) soluble in carbon dioxide at a pressure of between 600 and about 1050 pounds per square inch and at a temperature of between 5 and about 30 degrees Celsius; (2) specific gravity of greater than about 0.900 (the higher the specific gravity the better the organic solvent); (3) Hansen solubility parameters of about 13.5 – 18.0 (MPa)<sup>1/2</sup> for dispersion, about 4.0 – 6.5 (MPa)<sup>1/2</sup> for polar, and about 4.0 – 11.0 (MPa)<sup>1/2</sup> for hydrogen bonding.
- 10 14. *Ketones* exhibiting the following physical characteristics; (1) soluble in carbon dioxide at a pressure of between 600 and about 1050 pounds per square inch and at a temperature of between 5 and about 30 degrees Celsius; (2) specific gravity of greater than about 0.800 (the higher the specific gravity the better the organic solvent); (3) Hansen solubility parameters of about 13.0 – 19.0 (MPa)<sup>1/2</sup> for dispersion, about 3.0 – 8.0 (MPa)<sup>1/2</sup> for polar, and about 3.0 – 11.0 (MPa)<sup>1/2</sup> for hydrogen bonding.
- 15 15. *Aprotic solvents*. These include solvents that do not belong to any of the aforementioned solvent groups, contain no dissociable hydrogens, and exhibit the following physical characteristics; (1) soluble in carbon dioxide at a pressure of between 600 and about 1050 pounds per square inch and at a temperature of between 5 and about 30 degrees Celsius; (2) specific gravity of greater than about 0.900 (the higher the specific gravity the better the organic solvent); (3) Hansen solubility parameters of about 15.0 – 21.0 (MPa)<sup>1/2</sup> for dispersion, about 6.0 – 17.0 (MPa)<sup>1/2</sup> for polar, and about 4.0 – 13.0 (MPa)<sup>1/2</sup> for hydrogen bonding.
- 20 25

30 Preferably, in addition to the three physical properties described with respect to each above group, the organic solvent used in the present invention should also exhibit one or more of the following physical properties: (4) flash point greater than about 100 degrees Fahrenheit; and (5) evaporation rate of lower than about 50 (where n-butyl acetate=100). Most preferably, the organic solvent used in the present invention exhibits each of the foregoing characteristics (i.e., those identified as (1) through (5)).

The Hansen solubility parameters were developed to characterize solvents for the purpose of comparison. Each of the three parameters (i.e., dispersion, polar and hydrogen bonding) represents a different characteristic of solvency. In combination, the three parameters are a measure of the overall strength and selectivity of a solvent. The above Hansen solubility parameter ranges identify solvents that are good solvents for a wide range of substances and also exhibit a degree of solubility in liquid carbon dioxide. The Total Hansen solubility parameter, which is the square root of the sum of the squares of the three parameters mentioned previously, provides a more general description of the solvency of the organic solvents.

Any organic solvent or mixture of organic solvents from the groups specified and that meet at least properties 1 through 3, and preferably all 5 properties, is suitable for use in the present invention. Furthermore, the organic solvent should also have a low toxicity and a low environmental impact. Table 1 below shows the physical properties of a number of organic solvents that may be suitable for use in the present invention. In Table 1, the solvents are soluble in carbon dioxide between 570 psig/5°C and 830 psig/20°C.

TABLE 1								
Solvent	Soluble in carbon dioxide	Specific Gravity (20°C/20°C)	Flash Point (°F)	Evaporation Rate (n-butyl acetate = 100)	Hansen Solubility Parameters			
					Dispersion (MPa) <sup>1/2</sup>	Polar (MPa) <sup>1/2</sup>	Hydrogen Bonding (MPa) <sup>1/2</sup>	Total (MPa) <sup>1/2</sup>
Terpenes								
Pine Oil	y	.929 <sup>a</sup>	193 <sup>a</sup>	0.5 <sup>a</sup>	13.9 <sup>a</sup>	8.0 <sup>a</sup>	10.2 <sup>a</sup>	19.0 <sup>a</sup>
d-limonene	y	.843 <sup>c</sup> (25°C/25°C)	121 <sup>c</sup>	0.5 <sup>c</sup>	16.6 <sup>c</sup>	0.6 <sup>c</sup>	0.0 <sup>c</sup>	16.6 <sup>c</sup>
Halocarbons								
1,1,2-trifluoro- trichloroethane	y	1.57 <sup>b</sup>	none <sup>b</sup>	2100 <sup>b</sup>	14.7 <sup>b</sup>	1.6 <sup>b</sup>	0.0 <sup>b</sup>	14.7 <sup>b</sup>
n-propyl bromide	y	1.35 (25°C/25°C)	none	5.8	16.0 <sup>b</sup>	6.5 <sup>b</sup>	4.7 <sup>b</sup>	17.9
Perfluorohexane	y	1.67 <sup>i</sup>	none <sup>i</sup>	1000 <sup>d</sup>	12.1 <sup>d</sup>	0.0 <sup>d</sup>	0.0 <sup>d</sup>	12.1



TABLE 1								
Solvent	Soluble in carbon dioxide	Specific Gravity (20°C/20°C)	Flash Point (°F)	Evaporation Rate (n-butyl acetate = 100)	Hansen Solubility Parameters			
					Dispersion (MPa) <sup>1/2</sup>	Polar (MPa) <sup>1/2</sup>	Hydrogen Bonding (MPa) <sup>1/2</sup>	Total (MPa) <sup>1/2</sup>
Glycol Ethers								
Triethylene glycol mono-oleyl ether	y	0.92@ 15.5°C	>200 <sup>d</sup>	<1 <sup>d</sup>	13.3 <sup>a</sup>	3.1 <sup>a</sup>	8.4 <sup>a</sup>	16.0 <sup>a</sup>
Ethlan HB4*	y	1.12	>200 <sup>d</sup>	<0.5 <sup>d</sup>	17.4 <sup>d</sup>	9.2 <sup>d</sup>	13.0 <sup>d</sup>	23.6 <sup>d</sup>
Polyols								
Hexylene glycol	y	.921 <sup>b</sup>	201 <sup>b</sup>	1.0 <sup>b</sup>	15.8 <sup>b</sup>	8.4 <sup>b</sup>	17.8 <sup>b</sup>	25.2
Ethers								
Tetraethylene glycol dimethyl ether	y	1.005 <sup>b</sup>	285 <sup>b</sup>	~<0.5 <sup>d</sup>	15.7 <sup>b</sup>	2.0 <sup>b</sup>	8.2 <sup>b</sup>	17.8 <sup>b</sup>
Esters of Glycol Ethers								
Ethylene glycol diacetate	y	1.124 <sup>b</sup>	181 <sup>b</sup>	2.0 <sup>b</sup>	16.4 <sup>b</sup>	10.4 <sup>b</sup>	12.9 <sup>b</sup>	23.3 <sup>b</sup>
Esters of Carboxylic Acids								
Decyl acetates**	y	0.869 <sup>b</sup>	212 <sup>b</sup>	0.6 <sup>b</sup>	14.9 <sup>b</sup>	5.7 <sup>b</sup>	3.1 <sup>b</sup>	16.4 <sup>b</sup>
Tridecyl acetates***	y	0.875 <sup>b</sup>	261 <sup>b</sup>	0.1 <sup>b</sup>	15.1 <sup>b</sup>	5.1 <sup>b</sup>	1.6 <sup>b</sup>	16.1 <sup>b</sup>
Soy methyl esters*	y	0.87 <sup>c</sup> @ 25°C/25°C	425 <sup>c</sup>	< 0.5 <sup>c</sup>	16.1 <sup>c</sup>	4.9 <sup>c</sup>	5.9 <sup>c</sup>	17.8
Fatty Alcohols								
2-ethyl-hexanol	y	0.829 <sup>b</sup>	171 <sup>b</sup>	2.0 <sup>b</sup>	15.9 <sup>b</sup>	3.3 <sup>b</sup>	11.9 <sup>b</sup>	20.2 <sup>b</sup>
Aprotic Solvents								
Dimethylsulf-oxide	y	1.097 <sup>b</sup>	203 <sup>b</sup>	2.6 <sup>b</sup>	18.4 <sup>b</sup>	16.4 <sup>b</sup>	10.2 <sup>b</sup>	26.6 <sup>b</sup>
Dimethyl formamide	y	.94 <sup>b</sup>	136 <sup>b</sup>	20 <sup>b</sup>	17.4 <sup>b</sup>	13.7 <sup>b</sup>	11.2 <sup>b</sup>	24.7 <sup>b</sup>
Propylene carbonate	y	1.185 <sup>b</sup>	270 <sup>b</sup>	0.5 <sup>b</sup>	20.0 <sup>b</sup>	18.0 <sup>b</sup>	4.1 <sup>b</sup>	27.3 <sup>b</sup>

TABLE 1								
Solvent	Soluble In carbon dioxide	Specific Gravity (20°C/20°C)	Flash Point (°F)	Evaporation Rate (n-butyl acetate = 100)	Hansen Solubility Parameters			
					Dispersion (MPa) <sup>1/2</sup>	Polar (MPa) <sup>1/2</sup>	Hydrogen Bonding (MPa) <sup>1/2</sup>	Total (MPa) <sup>1/2</sup>
Siloxanes								
Octamethyl cyclotetra siloxane/deca methyl cyclopenta- siloxane++	y	0.96 <sup>a</sup> @ 25°C/25°C	144 <sup>a</sup>	<1 <sup>d</sup>	15.1 <sup>d</sup>	0.8 <sup>d</sup>	0.0 <sup>d</sup>	15.1 <sup>b</sup>
Hydrofluoroethers								
1-methoxy- nonafluoro- butane	y	1.52	none	900 <sup>d</sup>	13.7 <sup>d</sup>	6.1 <sup>d</sup>	8.2 <sup>d</sup>	17.1 <sup>d</sup>
Aliphatic Hydrocarbons								
Isoparaffins (DF 2000)	y	0.77	140	< 10	15.7 <sup>d</sup>	0.0 <sup>d</sup>	0.0 <sup>d</sup>	17.1 <sup>d</sup>
Dibasic Esters								
Dimethyl glutarate	y	1.084 <sup>b</sup>	225 <sup>b</sup>	<0.9 <sup>b</sup>	17.0 <sup>b</sup>	4.7 <sup>b</sup>	9.8 <sup>b</sup>	20.2 <sup>b</sup>

\*cc Phenyl - ω - hydroxy-poly (oxy 1,2 ethanediyl): Akzo Nobel

\*\* Exxate 1000; Exxon

\*\*\* Exxate 1300; Exxon

+ Soy Gold 1100; AG Environmental Products

++ SF 1204; General Electric Silicones

a Barton A.F.M.; Handbook of Solubility Parameters and Other Cohesion Parameters, 2<sup>nd</sup> Edition; CRC Press, 1991 (ISBN 0-8493-0176-9)

b Wypych, George; Handbook of Solvents, 2001; ChemTec (ISBN 1-895198-24-0)

c AG Environmental Products, website.

d Estimated.

e Clean Tech Proceedings 1998, pg 92

f Fluorochem USA

g GE Silicones Fluids Handbook, Bulletin No. 59 (9/91).

h Fedors Method: R.F. Fedors, Polymer Engineering and Science, 1974.

Referring now to FIG. 1, a block diagram of a cleaning system having separate vessels for cleaning and drying textiles is shown. The cleaning system 100 generally comprises a cleaning machine 102 having a cleaning vessel 110 operatively connected to, via one or more motor activated shafts (not shown), a perforated rotatable cleaning drum or wheel 112 within the cleaning vessel 110 with an inlet 114 to the cleaning vessel 110 and an outlet 116 from the cleaning vessel

110 through which cleaning fluids can pass. A drying machine 104 has a drying vessel 120 capable of being pressurized. The pressurizable drying vessel 120 is operatively connected to, via one or more motor activated shafts (not shown), a perforated rotatable drying drum or wheel 122 within the drying vessel 120 with an inlet 124 to the drying vessel 120 and an outlet 126 from the drying vessel 120 through which pressurized fluid solvent can pass. The cleaning vessel 110 and the drying vessel 120 can either be parts of the same machine, or they can comprise separate machines. Furthermore, both the cleaning and drying steps of this invention can be performed in the same vessel, as is described with respect to FIG. 2 below.

An organic solvent tank 130 holds any suitable organic solvent, as previously described, to be introduced to the cleaning vessel 110 through the inlet 114. A pressurized fluid solvent tank 132 holds pressurized fluid solvent to be added to the pressurizable drying vessel 120 through the inlet 124. Filtration assembly 140 contains one or more filters that continuously remove contaminants from the organic solvent from the cleaning vessel 110 as cleaning occurs.

The components of the cleaning system 100 are connected with lines 150-156, which transfer organic solvents and vaporized and pressurized fluid solvents between components of the system. The term "line" as used herein is understood to refer to a piping network or similar conduit capable of conveying fluid and, for certain purposes, is capable of being pressurized. The transfer of the organic solvents and vaporized and pressurized fluid solvents through the lines 150-156 is directed by valves 170-176 and pumps 190-193. While pumps 190-193 are shown in the described embodiment, any method of transferring liquid and/or vapor between components can be used, such as adding pressure to the component using a compressor to force the liquid and/or vapor from the component.

The textiles are cleaned with an organic solvent such as those previously described or mixtures thereof. The textiles may also be cleaned with a combination of organic solvent and pressurized fluid solvent, and this combination may be in varying proportions from about 50% by weight to 100% by weight of organic solvent and 0% by weight to 50% by weight of pressurized fluid solvent. In the cleaning process, the textiles are first sorted as necessary to place the textiles into groups suitable to be cleaned together. The textiles may then be spot treated as necessary to remove any stains that may not be removed during the cleaning process. The

textiles are then placed into the cleaning drum 112 of the cleaning system 100. It is preferred that the cleaning drum 112 be perforated to allow for free interchange of solvent between the cleaning drum 112 and the cleaning vessel 110 as well as to transport soil from the textiles to the filtration assembly 140.

5 After the textiles are placed in the cleaning drum 112, an organic solvent contained in the organic solvent tank 130 is added to the cleaning vessel 110 via line 152 by opening valve 171, closing valves 170, 172, 173 and 174, and activating pump 190 to pump organic solvent through the inlet 114 of the cleaning vessel 110. The organic solvent may contain one or more co-solvents, water, detergents, or  
10 other additives to enhance the cleaning capability of the cleaning system 100. Alternatively, one or more additives may be added directly to the cleaning vessel 110. Pressurized fluid solvent may also be added to the cleaning vessel 110 along with the organic solvent to enhance cleaning. Pressurized fluid solvent can be added to the cleaning vessel 110 via line 154 by opening valve 174, closing valves  
15 170, 171, 172, 173, and 175, and activating pump 192 to pump pressurized fluid solvent through the inlet 114 of the cleaning vessel 110. Of course, if pressurized fluid solvent is included in the cleaning cycle, the cleaning vessel 110 will need to be pressurized in the same manner as the drying vessel 120, as discussed below.

When a sufficient amount of the organic solvent, or combination of organic  
20 solvent and pressurized fluid solvent, is added to the cleaning vessel 110, the motor (not shown) is activated and the perforated cleaning drum 112 is agitated and/or rotated within cleaning vessel 110. During this phase, the organic solvent is continuously cycled through the filtration assembly 140 by opening valves 170 and 172, closing valves 171, 173 and 174, and activating pump 191. Filtration assembly  
25 140 may include one or more fine mesh filters to remove particulate contaminants from the organic solvent passing therethrough and may alternatively or in addition include one or more absorptive or adsorptive filters to remove water, dyes and other dissolved contaminants from the organic solvent. Exemplary configurations for filter assemblies that can be used to remove contaminants from either the organic solvent  
30 or the pressurized fluid solvent are described more fully in U.S. Application Serial No. 08/994,583 incorporated herein by reference. As a result, the organic solvent is pumped through outlet 116, valve 172, line 151, filter assembly 140, line 150, valve 170 and re-enters the cleaning vessel 110 via inlet 114. This cycling advantageously removes contaminants, including particulate contaminants and/or soluble

contaminants, from the organic solvent and reintroduces filtered organic solvent to the cleaning vessel 110 and agitating or rotating cleaning drum 112. Through this process, contaminants are removed from the textiles. Of course, in the event the cleaning vessel 110 is pressurized, this recirculation system will be maintained at the same pressure/temperature levels as those in cleaning vessel 110.

After sufficient time has passed so that the desired level of contaminants is removed from the textiles and organic solvent, the organic solvent is removed from the cleaning drum 112 and cleaning vessel 110 by opening valve 173, closing valves 170, 171, 172 and 174, and activating pump 191 to pump the organic solvent through outlet 116 via line 153. The cleaning drum 112 is then rotated at a high speed, such as 400-800 rpm, to further remove organic solvent from the textiles. The cleaning drum 112 is preferably perforated so that, when the textiles are rotated in the cleaning drum 112 at a high speed, the organic solvent can drain from the cleaning drum 112. Any organic solvent removed from the textiles by rotating the cleaning drum 112 at high speed is also removed from the cleaning drum 112 in the manner described above. After the organic solvent is removed from the cleaning drum 112, it can either be discarded or recovered and decontaminated for reuse using solvent recovery systems known in the art. Furthermore, multiple cleaning cycles can be used if desired, with each cleaning cycle using the same organic solvent or different organic solvents. If multiple cleaning cycles are used, each cleaning cycle can occur in the same cleaning vessel, or a separate cleaning vessel can be used for each cleaning cycle.

After a desired amount of the organic solvent is removed from the textiles by rotating the cleaning drum 112 at high speed, the textiles are moved from the cleaning drum 112 to the drying drum 122 within the drying vessel 120 in the same manner textiles are moved between machines in conventional cleaning systems. In an alternate embodiment, a single drum can be used in both the cleaning cycle and the drying cycle, so that, rather than transferring the textiles between the cleaning drum 112 and the drying drum 122, a single drum containing the textiles is transferred between the cleaning vessel 110 and the drying vessel 120. If the cleaning vessel 110 is pressurized during the cleaning cycle, it must be depressurized before the textiles are removed. Once the textiles have been placed in the drying drum 122, pressurized fluid solvent, such as that contained in the carbon dioxide tank 132, is added to the drying vessel 120 via lines 154 and 155 by

opening valve 175, closing valves 174 and 176, and activating pump 192 to pump pressurized fluid solvent through the inlet 124 of the drying vessel 120 via lines 154 and 155. When pressurized fluid solvent is added to the drying vessel 120, the organic solvent remaining on the textiles dissolves in the pressurized fluid solvent.

5 After a sufficient amount of pressurized fluid solvent is added so that the desired level of organic solvent has been dissolved, the pressurized fluid solvent and organic solvent combination is removed from the drying vessel 120, and therefore also from the drying drum 122, by opening valve 176, closing valve 175 and activating pump 193 to pump the pressurized fluid solvent and organic solvent combination through outlet 126 via line 156. If desired, this process may be  
10 repeated to remove additional organic solvent. The drying drum 122 is then rotated at a high speed, such as 150-350 rpm, to further remove the pressurized fluid solvent and organic solvent combination from the textiles. The drying drum 122 is preferably perforated so that, when the textiles are rotated in the drying drum 122 at  
15 a high speed, the pressurized fluid solvent and organic solvent combination can drain from the drying drum 122. Any pressurized fluid solvent and organic solvent combination removed from the textiles by spinning the drying drum 122 at high speed is also pumped from the drying vessel 120 in the manner described above. After the pressurized fluid solvent and organic solvent combination is removed from  
20 the drying vessel 120, it can either be discarded or separated and recovered for reuse with solvent recovery systems known in the art. Note that, while preferred, it is not necessary to include a high speed spin cycle to remove pressurized fluid solvent from the textiles.

After a desired amount of the pressurized fluid solvent is removed from the  
25 textiles by rotating the drying drum 122, the drying vessel 120 is depressurized over a period of about 5-15 minutes. The depressurization of the drying vessel 120 vaporizes any remaining pressurized fluid solvent, leaving dry, solvent-free textiles in the drying drum 122. The pressurized fluid solvent that has been vaporized is then removed from the drying vessel 120 by opening valve 176, closing valve 175, and  
30 activating pump 193. As a result, the vaporized pressurized fluid solvent is pumped through the outlet 126, line 156 and valve 176, where it can then either be vented to the atmosphere or recovered and recompressed for reuse.

While the cleaning system 100 has been described as a complete system, an existing conventional dry cleaning system may be converted for use in accordance

with the present invention. To convert a conventional dry cleaning system, the organic solvent described above is used to clean textiles in the conventional system. A separate pressurized vessel is added to the conventional system for drying the textiles with pressurized fluid solvent. Thus, the conventional system is converted for use with a pressurized fluid solvent. For example, the system in FIG. 1 could represent such a converted system, wherein the components of the cleaning machine 102 are conventional, and the pressurized fluid solvent tank 132 is not in communication with the cleaning vessel 100. In such a situation, the drying machine 104 is the add-on part of the conventional cleaning machine.

Furthermore, while the system shown in FIG. 1 comprises a single cleaning vessel, multiple cleaning vessels could be used, so that the textiles are subjected to multiple cleaning steps, with each cleaning step carried out in a different cleaning vessel using the same or different organic solvents in each step. The description of the single cleaning vessel is merely for purposes of description and should not be construed as limiting the scope of the invention.

Referring now to FIG. 2, a block diagram of an alternate embodiment of the present invention, a cleaning system having a single chamber for cleaning and drying the textiles, is shown. The cleaning system 200 generally comprises a cleaning machine having a pressurizable vessel 210. The vessel 210 is operatively connected to, via one or more motor activated shafts (not shown), a perforated rotatable drum or wheel 212 within the vessel 210 with an inlet 214 to the vessel 210 and an outlet 216 from the vessel 210 through which dry cleaning fluids can pass.

An organic solvent tank 220 holds any suitable organic solvent, such as those described above, to be introduced to the vessel 210 through the inlet 214. A pressurized fluid solvent tank 222 holds pressurized fluid solvent to be added to the vessel 210 through the inlet 214. Filtration assembly 224 contains one or more filters that continuously remove contaminants from the organic solvent from the vessel 210 and drum 212 as cleaning occurs.

The components of the cleaning system 200 are connected with lines 230-234 that transfer organic solvents and vaporized and pressurized fluid solvent between components of the system. The term "line" as used herein is understood to refer to a piping network or similar conduit capable of conveying fluid and, for certain purposes, is capable of being pressurized. The transfer of the organic solvents and vaporized and pressurized fluid solvent through the lines 230-234 is directed by

valves 250-254 and pumps 240-242. While pumps 240-242 are shown in the described embodiment, any method of transferring liquid and/or vapor between components can be used, such as adding pressure to the component using a compressor to force the liquid and/or vapor from the component.

5       The textiles are cleaned with an organic solvent such as those previously described. The textiles may also be cleaned with a combination of organic solvent and pressurized fluid solvent, and this combination may be in varying proportions of 50-100% by weight organic solvent and 0-50% by weight pressurized fluid solvent. In the cleaning process, the textiles are first sorted as necessary to place the textiles  
10       into groups suitable to be cleaned together. The textiles may then be spot treated as necessary to remove any stains that may not be removed during the cleaning process. The textiles are then placed into the drum 212 within the vessel 210 of the cleaning system 200. It is preferred that the drum 212 be perforated to allow for free interchange of solvent between the drum 212 and the vessel 210 as well as to  
15       transport soil from the textiles to the filtration assembly 224.

After the textiles are placed in the drum 212, an organic solvent contained in the organic solvent tank 220 is added to the vessel 210 via line 231 by opening valve 251, closing valves 250, 252, 253 and 254, and activating pump 242 to pump  
20       organic solvent through the inlet 214 of the vessel 210. The organic solvent may contain one or more co-solvents, detergents, water, or other additives to enhance the cleaning capability of the cleaning system 200 or other additives to impart other desirable attributes to the articles being treated. Alternatively, one or more additives may be added directly to the vessel. Pressurized fluid solvent may also be added to the vessel 210 along with the organic solvent to enhance cleaning. The pressurized  
25       fluid solvent is added to the vessel 210 via line 230 by opening valve 250, closing valves 251, 252, 253 and 254, and activating pump 240 to pump the pressurized fluid solvent through the inlet 214 of the vessel 210.

When the desired amount of the organic solvent, or combination of organic solvent and pressurized fluid solvent as described above, is added to the vessel 210,  
30       the motor (not shown) is activated and the drum 212 is agitated and/or rotated. During this phase, the organic solvent, as well as pressurized fluid solvent if used in combination, is continuously cycled through the filtration assembly 224 by opening valves 252 and 253, closing valves 250, 251 and 254, and activating pump 241. Filtration assembly 224 may include one or more fine mesh filters to remove



particulate contaminants from the organic solvent and pressurized fluid solvent passing therethrough and may alternatively or in addition include one or more absorptive or adsorptive filters to remove water, dyes, and other dissolved contaminants from the organic solvent. Exemplary configurations for filter assemblies that can be used to remove contaminants from either the organic solvent or the pressurized fluid solvent are described more fully in U.S. Application Serial No. 08/994,583 incorporated herein by reference. As a result, the organic solvent is pumped through outlet 216, valve 253, line 233, filter assembly 224, line 232, valve 252 and reenters the vessel 210 via inlet 214. This cycling advantageously removes contaminants, including particulate contaminants and/or soluble contaminants, from the organic solvent and pressurized fluid solvent and reintroduces filtered solvent to the vessel 210. Through this process, contaminants are removed from the textiles.

After sufficient time has passed so that the desired level of contaminants is removed from the textiles and solvents, the organic solvent is removed from the vessel 210 and drum 212 by opening valve 254, closing valves 250, 251, 252 and 253, and activating pump 241 to pump the organic solvent through outlet 216 and line 234. If pressurized fluid solvent is used in combination with organic solvent, it may be necessary to first separate the pressurized fluid solvent from the organic solvent. The organic solvent can then either be discarded or, preferably, contaminants may be removed from the organic solvent and the organic solvent recovered for further use. Contaminants may be removed from the organic solvent with solvent recovery systems known in the art. The drum 212 is then rotated at a high speed, such as 400-800 rpm, to further remove organic solvent from the textiles. The drum 212 is preferably perforated so that, when the textiles are rotated in the drum 212 at a high speed, the organic solvent can drain from the cleaning drum 212. Any organic solvent removed from the textiles by rotating the drum 212 at high speed can also either be discarded or recovered for further use.

After a desired amount of organic solvent is removed from the textiles by rotating the drum 212, pressurized fluid solvent contained in the pressurized fluid tank 222 is added to the vessel 210 by opening valve 250, closing valves 251, 252, 253 and 254, and activating pump 240 to pump pressurized fluid solvent through the inlet 214 of the pressurizable vessel 210 via line 230. When pressurized fluid solvent is added to the vessel 210, organic solvent remaining on the textiles dissolves in the pressurized fluid solvent.

After a sufficient amount of pressurized fluid solvent is added so that the desired level of organic solvent has been dissolved, the pressurized fluid solvent and organic solvent combination is removed from the vessel 210 by opening valve 254, closing valves 250, 251, 252 and 253, and activating pump 241 to pump the pressurized fluid solvent and organic solvent combination through outlet 216 and line 234. Note that pump 241 may actually require two pumps, one for pumping the low pressure organic solvent in the cleaning cycle and one for pumping the pressurized fluid solvent in the drying cycle.

The pressurized fluid solvent and organic solvent combination can then either be discarded or the combination may be separated and the organic solvent and pressurized fluid solvent separately recovered for further use. The drum 212 is then rotated at a high speed, such as 150-350 rpm, to further remove pressurized fluid solvent and organic solvent combination from the textiles. Any pressurized fluid solvent and organic solvent combination removed from the textiles by spinning the drum 212 at high speed can also either be discarded or retained for further use. Note that, while preferred, it is not necessary to include a high speed spin cycle to remove pressurized fluid solvent from the textiles.

After a desired amount of the pressurized fluid solvent is removed from the textiles by rotating the drum 212, the vessel 210 is depressurized over a period of about 5-15 minutes. The depressurization of the vessel 210 vaporizes the pressurized fluid solvent, leaving dry, solvent-free textiles in the drum 212. The pressurized fluid solvent that has been vaporized is then removed from the vessel 210 by opening valve 254, closing valves 250, 251, 252 and 253, and activating pump 241 to pump the vaporized pressurized fluid solvent through outlet 216 and line 234. Note that while a single pump is shown as pump 241, separate pumps may be necessary to pump organic solvent, pressurized fluid solvent and pressurized fluid solvent vapors, at pump 241. The remaining vaporized pressurized fluid solvent can then either be vented into the atmosphere or compressed back into pressurized fluid solvent for further use.

As discussed above, terpenes, halohydrocarbons, certain glycol ethers, polyols, ethers, esters of glycol ethers, esters of fatty acids and other long chain carboxylic acids, fatty alcohols and other long-chain alcohols, short-chain alcohols, polar aprotic solvents, cyclic methyl siloxanes, hydrofluoroethers, dibasic esters, and aliphatic hydrocarbons solvents or similar solvents or mixtures of such solvents are

organic solvents that can be used in the present invention, as shown in the test results below. Table 2 shows results of detergency testing for each of a number of solvents that may be suitable for use in the present invention. Table 3 shows results of testing of drying and extraction of those solvents using densified carbon dioxide.

5 Detergency tests were performed using a number of different solvents without detergents, co-solvents, or other additives. The solvents selected for testing include organic solvents and liquid carbon dioxide. Two aspects of detergency were investigated - soil removal and soil redeposition. The former refers to the ability of a solvent to remove soil from a substrate while the latter refers to the ability of a solvent to prevent soil from being redeposited on a substrate during the cleaning process. Wascherei Forschungs Institute, Krefeld Germany ("WFK") standard soiled  
10 swatches that have been stained with a range of insoluble materials and WFK white cotton swatches, both obtained from TESTFABRICS, Inc., were used to evaluate soil removal and soil redeposition, respectively.

15 Soil removal and redeposition for each solvent was quantified using the Delta Whiteness Index. This method entails measuring the Whiteness Index of each swatch before and after processing. The Delta Whiteness Index is calculated by subtracting the Whiteness Index of the swatch before processing from the Whiteness Index of the swatch after processing. The Whiteness Index is a function of the light reflectance of the swatch and in this application is an indication of the amount of soil  
20 on the swatch. More soil results in a lower light reflectance and Whiteness Index for the swatch. The Whiteness indices were measured using a reflectometer manufactured by Hunter Laboratories.

25 Organic solvent testing was carried out in a Launder-Ometer while the densified carbon dioxide testing was carried out in a Parr Bomb. After measuring their Whiteness Indices, two WFK standard soil swatches and two WFK white cotton swatches were placed in a Launder-Ometer cup with 25 stainless steel ball bearings and 150 mL of the solvent of interest. The cup was then sealed, placed in the Launder-Ometer and agitated for a specified length of time. Afterwards, the  
30 swatches were removed and placed in a Parr Bomb equipped with a mesh basket. Approximately 1.5 liters of liquid carbon dioxide between 5°C and 25°C and 570 psig and 830 psig was transferred to the Parr Bomb. After several minutes the Parr Bomb was vented and the dry swatches removed and allowed to reach room temperature. Testing of densified carbon dioxide was carried out in the same

manner but test swatches were treated for 20 minutes. During this time the liquid carbon dioxide was stirred using an agitator mounted on the inside cover of the Parr bomb. The Whiteness Index of the processed swatches was determined using the reflectometer. The two Delta Whiteness Indices obtained for each pair of swatches were averaged. The results are presented in Table 2.

Because the Delta Whiteness Index is calculated by subtracting the Whiteness Index of a swatch before processing from the Whiteness Index value after processing, a positive Delta Whiteness Index indicates that there was an increase in Whiteness Index as a result of processing. In practical terms, this means that soil was removed during processing. In fact, the higher the Delta Whiteness Value, the more soil was removed from the swatch during processing. Each of the organic solvents tested exhibited soil removal capabilities. The WFK white cotton swatches exhibited a decrease in Delta Whiteness Indices indicating that the soil was deposited on the swatches during the cleaning process. Therefore, a "less negative" Delta Whiteness Index suggests that less soil was redeposited.

TABLE 2

Solvent	Cleaning Time (min.)	Delta Whiteness Values	
		Insoluble Soil Removal	Insoluble Soil Redeposition
Liquid carbon dioxide (neat)	20	3.36	-1.23
Pine oil	12	8.49	-6.84
d-limonene	12	10.6	-9.2
1,1,2 trichlorotrifluoroethane	12	11.7	-14.48
N-propyl bromide	12	11.18	-9.45
Perfluorohexane	12	2.09	-3.42
Triethylene glycol mono-oleyl ether (Volpo 3)	12	10.54*	-1.86*
$\alpha$ -phenyl - $\omega$ -hydroxy-poly (oxy-1,2-ethanediyl)	12	1.54**	-13.6**
Hexylene glycol	12	6.9	-1.4
Tetraethylene glycol dimethyl ether	12	10.08	-4.94
Ethylene glycol diacetate	12	6.29	-3.39
Decyl acetates (Exxate 1000)	12	11.69	-8.6
Tridecyl acetates (Exxate 1300)	12	11.24	-4.88
Soy methyl esters (SoyGold 1100)	12	5.81	-7.71
2-ethylhexanol	12	12.6	-3.4
Propylene carbonate	12	2.99	-1.82
Dimethylsulfoxide	12	5.84	-0.22
Dimethylformamide	12	7.24	-10.09
Isoparaffins (DF-2000)	12	11.23	-5.95
Dimethyl glutarate	12	9.04	-1.23

\* After two extraction cycles

\*\* After three extraction cycles.

5 To evaluate the ability of densified carbon dioxide to extract organic solvent from a substrate, WFK white cotton swatches were used. One swatch was weighed dry and then immersed in an organic solvent sample. Excess solvent was removed from the swatch using a ringer manufactured by Atlas Electric Devices Company. The damp swatch was re-weighed to determine the amount of solvent retained in the fabric. After placing the damp swatch in a Parr Bomb densified carbon dioxide was transferred to the Parr Bomb. The temperature and pressure of the densified carbon dioxide for all of the trials ranged from 5°C to 20°C and from 570 psig - 830 psig. After five minutes the Parr Bomb was vented and the swatch removed. The swatch was next subjected to Soxhlet extraction using methylene chloride for a minimum of two hours. This apparatus enables the swatch to be continuously extracted to remove the organic solvent from the swatch. After determining the concentration of

10

15

the organic solvent in the extract using gas chromatography, the amount of organic solvent remaining on the swatch after exposure to densified carbon dioxide was calculated by multiplying the concentration of the organic solvent in the extract by the volume of the extract. A different swatch was used for each of the tests. The results of these tests are included in Table 3. As the results indicate, the extraction process using densified carbon dioxide is extremely effective.

TABLE 3

Solvent	Weight of Solvent on Test Swatch (grams)		Percentage by Weight of Solvent Removed from Swatch
	Before Extraction	After Extraction	
Pine oil	7.8	0.1835	97.66%
d-Limonene	5.8	0.0014	99.98%
1,1,2-Trichlorotrifluoroethane	1.4	0.0005	99.96%
n-Propyl bromide	2.8	<0.447	>84%
Perfluorohexane	1.0	0.0006	99.94%
Triethylene glycol monooleyl ether(7)	0.8	0.1824	77.88%
$\alpha$ -phenyl - $\omega$ - hydroxy-poly(oxy 1,2-ethanediyl); (Ethylan HB4)	16.0	5.7	64.5%
Hexylene glycol	4.9	0.3481	92.87%
Tetraethylene glycol dimethyl ether	5.2	.1310	97.48%
Ethylene glycol diacetate	5.3	0.0418	99.21%
Decyl acetate(2)	2.4	0.0015	99.94%
Tridecyl acetate(1)	4.8	0.0605	98.75%
Soy methyl esters (8)	4.9	0.0720	98.54%
2-Ethylhexanol	0.5	0.0599	99.09%
Propylene carbonate	6.6	0.0599	99.09%
Dimethyl sulfoxide	3.3	0.5643	82.69%
Dimethylformamide	3.0	0.0635	97.88%
Octamethylcyclotetrasiloxane/Decamethylcyclopentasiloxane(4)	5.5	0.0017	99.97%
1-Methoxynonofluorobutane (6)	0.7	not detected	~ 100%
Isoparaffins (5)	4.3	0.0019	99.96%
Dimethyl glutarate(3)†	5.8	0.0090	99.85%

Notes on Table 3: (1) Exxate 1300 (Exxon); (2) Exxate 1000 (Exxon); (3) DBE-5 (DuPont); (4) SF1204 (General Electric Silicones); (5) DF-2000 (Exxon); (6) HFE-7100 (3M); (7) Volpo 3 (Croda); (8) Soy Gold 1100 (AG Environmental Products)

It is to be understood that a wide range of changes and modifications to the embodiments described above will be apparent to those skilled in the art and are contemplated. It is, therefore, intended that the foregoing detailed description be regarded as illustrative rather than limiting, and that it be understood that it is the following claims, including all equivalents, that are intended to define the spirit and scope of the invention.

## CLAIMS

What is claimed is:

1. A process for cleaning substrates comprising:  
placing the substrates to be cleaned in a vessel;  
adding organic solvent to the vessel;  
cleaning the substrates with an organic solvent;  
removing a portion of the organic solvent from the vessel;  
adding pressurized fluid solvent to the vessel;  
removing the pressurized fluid solvent from the vessel; and  
removing the substrates from the vessel.

2. The process of claim 1 wherein the organic solvent comprises a cyclic terpene.

3. The process of claim 2 wherein the cyclic terpene:  
is soluble in carbon dioxide between 600 and 1050 pounds per square inch and between 5 and 30 degrees Celsius;  
has a specific gravity of greater than approximately 0.800;  
has a dispersion Hansen solubility parameter of between  $13.0 \text{ (MPa)}^{1/2}$  and  $17.5 \text{ (MPa)}^{1/2}$ ;  
has a polar Hansen solubility parameter of between  $0.5 \text{ (MPa)}^{1/2}$  and  $9.0 \text{ (MPa)}^{1/2}$ ; and  
has a hydrogen bonding Hansen solubility parameter of between  $0.0 \text{ (MPa)}^{1/2}$  and  $10.5 \text{ (MPa)}^{1/2}$ .

4. The process of claim 3 wherein the cyclic terpene further:  
has an evaporation rate of lower than 50 (based on n-butyl acetate = 100); and  
has a flash point greater than 100 degrees Fahrenheit.

5. The process of claim 4 wherein the cyclic terpene is selected from a group including  $\alpha$ -terpene isomers; pine oil;  $\alpha$ -pinene isomers; d-limonene; and mixtures thereof.

6. The process of claim 1 wherein the organic solvent comprises a halocarbon.

7. The process of claim 6 wherein the halocarbon:

is soluble in carbon dioxide between 600 and 1050 pounds per square inch and between 5 and 30 degrees Celsius;

has a specific gravity of greater than approximately 1.100;

has a dispersion Hansen solubility parameter of between  $10.0 \text{ (MPa)}^{1/2}$  and  $17.0 \text{ (MPa)}^{1/2}$ ;

has a polar Hansen solubility parameter of between  $0.0 \text{ (MPa)}^{1/2}$  and  $7.0 \text{ (MPa)}^{1/2}$ ; and

has a hydrogen bonding Hansen solubility parameter of between  $0.0 \text{ (MPa)}^{1/2}$  and  $5.0 \text{ (MPa)}^{1/2}$ .

8. The process of claim 7 wherein the halocarbon further:

has an evaporation rate of lower than 50 (based on n-butyl acetate = 100); and

has a flash point greater than 100 degrees Fahrenheit.

9. The process of claim 8 wherein the halocarbon is selected from a group including chlorinated hydrocarbons; fluorinated hydrocarbons; brominated hydrocarbons; and mixtures thereof.

10. The process of claim 1 wherein the organic solvent comprises a glycol ether.

11. The process of claim 10 wherein the glycol ether:

is soluble in carbon dioxide between 600 and 1050 pounds per square inch and between 5 and 30 degrees Celsius;

has a specific gravity of greater than approximately 0.800;

has a dispersion Hansen solubility parameter of between  $13.0 \text{ (MPa)}^{1/2}$  and  $19.5 \text{ (MPa)}^{1/2}$ ;

has a polar Hansen solubility parameter of between  $3.0 \text{ (MPa)}^{1/2}$  and  $7.5 \text{ (MPa)}^{1/2}$ ; and



has a hydrogen bonding Hansen solubility parameter of between 8.0 (MPa)<sup>1/2</sup> and 17.0 (MPa)<sup>1/2</sup>.

12. The process of claim 11 wherein the glycol ether further:  
5 has an evaporation rate of lower than 50 (based on n-butyl acetate = 100); and  
has a flash point greater than 100 degrees Fahrenheit.

13. The process of claim 12 wherein the glycol ether is selected from a  
10 group including monoethylene glycol ether; diethylene glycol ether; triethylene glycol ether; monopropylene glycol ether; dipropylene glycol ether; tripropylene glycol ether; and mixtures thereof.

14. The process of claim 1 wherein the organic solvent comprises a polyol.  
15

15. The process of claim 14 wherein the polyol:  
is soluble in carbon dioxide between 600 and 1050 pounds per square inch and between 5 and 30 degrees Celsius;  
has a specific gravity of greater than approximately 0.920;  
20 has a dispersion Hansen solubility parameter of between 14.0 (MPa)<sup>1/2</sup> and 18.2 (MPa)<sup>1/2</sup>;  
has a polar Hansen solubility parameter of between 4.5 (MPa)<sup>1/2</sup> and 20.5 (MPa)<sup>1/2</sup>; and  
has a hydrogen bonding Hansen solubility parameter of between 15.0 (MPa)<sup>1/2</sup> and 30.0 (MPa)<sup>1/2</sup>.  
25

16. The process of claim 15 wherein the polyol further:  
has an evaporation rate of lower than 50 (based on n-butyl acetate = 100); and  
30 has a flash point greater than 100 degrees Fahrenheit.

17. The process of claim 16 wherein the polyol contains two or more hydroxyl radicals.

18. The process of claim 1 wherein the organic solvent comprises an ether.

19. The process of claim 18 wherein the ether:

is soluble in carbon dioxide between 600 and 1050 pounds per square inch and between 5 and 30 degrees Celsius;

has a specific gravity of greater than approximately 0.800;

has a dispersion Hansen solubility parameter of between  $14.5 \text{ (MPa)}^{1/2}$  and  $20.0 \text{ (MPa)}^{1/2}$ ;

has a polar Hansen solubility parameter of between  $1.5 \text{ (MPa)}^{1/2}$  and  $6.5 \text{ (MPa)}^{1/2}$ ; and

has a hydrogen bonding Hansen solubility parameter of between  $5.0 \text{ (MPa)}^{1/2}$  and  $10.0 \text{ (MPa)}^{1/2}$ .

20. The process of claim 19 wherein the ether further:

has an evaporation rate of lower than 50 (based on n-butyl acetate =

100); and

has a flash point greater than 100 degrees Fahrenheit.

21. The process of claim 20 wherein the ether contains no free hydroxyl radicals.

22. The process of claim 1 wherein the organic solvent comprises an ester of glycol ethers.

23. The process of claim 22 wherein the ester of glycol ethers:

is soluble in carbon dioxide between 600 and 1050 pounds per square inch and between 5 and 30 degrees Celsius;

has a specific gravity of greater than approximately 0.800;

has a dispersion Hansen solubility parameter of between  $15.0 \text{ (MPa)}^{1/2}$  and  $20.0 \text{ (MPa)}^{1/2}$ ;

has a polar Hansen solubility parameter of between  $3.0 \text{ (MPa)}^{1/2}$  and  $10.0 \text{ (MPa)}^{1/2}$ ; and

has a hydrogen bonding Hansen solubility parameter of between  $8.0 \text{ (MPa)}^{1/2}$  and  $16.0 \text{ (MPa)}^{1/2}$ .

24. The process of claim 23 wherein the ester of glycol ethers further:  
has an evaporation rate of lower than 50 (based on n-butyl acetate =  
100); and  
has a flash point greater than 100 degrees Fahrenheit.

25. The process of claim 1 wherein the organic solvent comprises an ester  
of monobasic carboxylic acids.

26. The process of claim 25 wherein the ester of monobasic carboxylic  
acids:

is soluble in carbon dioxide between 600 and 1050 pounds per square  
inch and between 5 and 30 degrees Celsius;

has a specific gravity of greater than approximately 0.800;

has a dispersion Hansen solubility parameter of between  $13.0 \text{ (MPa)}^{1/2}$   
and  $17.0 \text{ (MPa)}^{1/2}$ ;

has a polar Hansen solubility parameter of between  $2.0 \text{ (MPa)}^{1/2}$  and  
 $7.5 \text{ (MPa)}^{1/2}$ ; and

has a hydrogen bonding Hansen solubility parameter of between  $1.5 \text{ (MPa)}^{1/2}$   
and  $6.5 \text{ (MPa)}^{1/2}$ .

27. The process of claim 26 wherein the ester of monobasic carboxylic  
acids further:

has an evaporation rate of lower than 50 (based on n-butyl acetate =

100); and

has a flash point greater than 100 degrees Fahrenheit.

28. The process of claim 1 wherein the organic solvent comprises a fatty  
alcohol.

29. The process of claim 28 wherein the fatty alcohol:

is soluble in carbon dioxide between 600 and 1050 pounds per square  
inch and between 5 and 30 degrees Celsius;

has a specific gravity of greater than approximately 0.800;

34

has a dispersion Hansen solubility parameter of between  $13.3 \text{ (MPa)}^{1/2}$  and  $18.4 \text{ (MPa)}^{1/2}$ ;

has a polar Hansen solubility parameter of between  $3.1 \text{ (MPa)}^{1/2}$  and  $18.8 \text{ (MPa)}^{1/2}$ ; and

has a hydrogen bonding Hansen solubility parameter of between  $8.4 \text{ (MPa)}^{1/2}$  and  $22.3 \text{ (MPa)}^{1/2}$ .

30. The process of claim 29 wherein the fatty alcohol further:  
has an evaporation rate of lower than 50 (based on n-butyl acetate = 100); and  
has a flash point greater than 100 degrees Fahrenheit.

31. The process of claim 30 wherein, in the fatty alcohol, the carbon chain adjacent to the hydroxyl group contains at least five carbon atoms.

32. The process of claim 1 wherein the organic solvent comprises a short chain alcohol.

33. The process of claim 32 wherein the short chain alcohol:  
is soluble in carbon dioxide between 600 and 1050 pounds per square inch and between 5 and 30 degrees Celsius;  
has a specific gravity of greater than approximately 0.800;  
has a dispersion Hansen solubility parameter of between  $13.5 \text{ (MPa)}^{1/2}$  and  $18.0 \text{ (MPa)}^{1/2}$ ;

has a polar Hansen solubility parameter of between  $3.0 \text{ (MPa)}^{1/2}$  and  $9.0 \text{ (MPa)}^{1/2}$ ; and  
has a hydrogen bonding Hansen solubility parameter of between  $9.0 \text{ (MPa)}^{1/2}$  and  $16.5 \text{ (MPa)}^{1/2}$ .

34. The process of claim 33 wherein the short chain alcohol further:  
has an evaporation rate of lower than 50 (based on n-butyl acetate = 100); and  
has a flash point greater than 100 degrees Fahrenheit.

35. The process of claim 34 wherein, in the short chain alcohol, the carbon chain adjacent to the hydroxyl group contains no more than four carbon atoms.

36. The process of claim 1 wherein the organic solvent comprises a siloxane.

37. The process of claim 36 wherein the siloxane:

is soluble in carbon dioxide between 600 and 1050 pounds per square inch and between 5 and 30 degrees Celsius;

has a specific gravity of greater than approximately 0.900;

has a dispersion Hansen solubility parameter of between  $14.0 \text{ (MPa)}^{1/2}$  and  $18.0 \text{ (MPa)}^{1/2}$ ;

has a polar Hansen solubility parameter of between  $0.0 \text{ (MPa)}^{1/2}$  and  $4.5 \text{ (MPa)}^{1/2}$ ; and

has a hydrogen bonding Hansen solubility parameter of between  $0.0 \text{ (MPa)}^{1/2}$  and  $4.5 \text{ (MPa)}^{1/2}$ .

38. The process of claim 37 wherein the siloxane:

has an evaporation rate of lower than 50 (based on n-butyl acetate =

100); and

has a flash point greater than 100 degrees Fahrenheit.

39. The process of claim 1 wherein the organic solvent comprises a hydrofluoroether.

40. The process of claim 39 wherein the hydrofluoroether:

is soluble in carbon dioxide between 600 and 1050 pounds per square inch and between 5 and 30 degrees Celsius;

has a specific gravity of greater than approximately 1.500;

has a dispersion Hansen solubility parameter of between  $12.0 \text{ (MPa)}^{1/2}$  and  $18.0 \text{ (MPa)}^{1/2}$ ;

has a polar Hansen solubility parameter of between  $4.0 \text{ (MPa)}^{1/2}$  and  $10.0 \text{ (MPa)}^{1/2}$ ; and

has a hydrogen bonding Hansen solubility parameter of between 1.5 (MPa)<sup>1/2</sup> and 9.0 (MPa)<sup>1/2</sup>.

41. The process of claim 40 wherein the hydrofluoroether:  
5 has an evaporation rate of lower than 50 (based on n-butyl acetate = 100); and  
has a flash point greater than 100 degrees Fahrenheit.

42. The process of claim 1 wherein the organic solvent comprises an  
10 aliphatic hydrocarbon.

43. The process of claim 42 wherein the aliphatic hydrocarbon:  
is soluble in carbon dioxide between 600 and 1050 pounds per square  
inch and between 5 and 30 degrees Celsius;  
15 has a specific gravity of greater than approximately 0.700;  
has a dispersion Hansen solubility parameter of between 14.0 (MPa)<sup>1/2</sup>  
and 17.0 (MPa)<sup>1/2</sup>;  
has a polar Hansen solubility parameter of between 0.0 (MPa)<sup>1/2</sup> and  
2.0 (MPa)<sup>1/2</sup>; and  
20 has a hydrogen bonding Hansen solubility parameter of between 0.0  
(MPa)<sup>1/2</sup> and 2.0 (MPa)<sup>1/2</sup>.

44. The process of claim 43 wherein the aliphatic hydrocarbon:  
has an evaporation rate of lower than 50 (based on n-butyl acetate =  
25 100); and  
has a flash point greater than 100 degrees Fahrenheit.

45. The process of claim 1 wherein the organic solvent comprises an ester  
of dibasic carboxylic acids.

30 46. The process of claim 45 wherein the ester of dibasic carboxylic acids:  
is soluble in carbon dioxide between 600 and 1050 pounds per square  
inch and between 5 and 30 degrees Celsius;  
has a specific gravity of greater than approximately 0.900;

has a dispersion Hansen solubility parameter of between  $13.5 \text{ (MPa)}^{1/2}$  and  $18.0 \text{ (MPa)}^{1/2}$ ;

has a polar Hansen solubility parameter of between  $4.0 \text{ (MPa)}^{1/2}$  and  $6.5 \text{ (MPa)}^{1/2}$ ; and

has a hydrogen bonding Hansen solubility parameter of between  $4.0 \text{ (MPa)}^{1/2}$  and  $11.0 \text{ (MPa)}^{1/2}$ .

47. The process of claim 46 wherein the ester of dibasic carboxylic acids:  
has an evaporation rate of lower than 50 (based on n-butyl acetate = 100); and  
has a flash point greater than 100 degrees Fahrenheit.

48. The process of claim 1 wherein the organic solvent comprises a ketone.

49. The process of claim 48 wherein the ketone:  
is soluble in carbon dioxide between 600 and 1050 pounds per square inch and between 5 and 30 degrees Celsius;  
has a specific gravity of greater than approximately 0.800;  
has a dispersion Hansen solubility parameter of between  $13.0 \text{ (MPa)}^{1/2}$  and  $19.0 \text{ (MPa)}^{1/2}$ ;  
has a polar Hansen solubility parameter of between  $3.0 \text{ (MPa)}^{1/2}$  and  $8.0 \text{ (MPa)}^{1/2}$ ; and  
has a hydrogen bonding Hansen solubility parameter of between  $3.0 \text{ (MPa)}^{1/2}$  and  $11.0 \text{ (MPa)}^{1/2}$ .

50. The process of claim 49 wherein the ketone:  
has an evaporation rate of lower than 50 (based on n-butyl acetate = 100); and  
has a flash point greater than 100 degrees Fahrenheit.

51. The process of claim 1 wherein the organic solvent comprises an aprotic solvent that contains no dissociable hydrogens.

52. The process of claim 51 wherein the aprotic solvent:  
is soluble in carbon dioxide between 600 and 1050 pounds per square  
inch and between 5 and 30 degrees Celsius;

has a specific gravity of greater than approximately 0.900;

5 has a dispersion Hansen solubility parameter of between  $15.0 \text{ (MPa)}^{1/2}$   
and  $21.0 \text{ (MPa)}^{1/2}$ ;

has a polar Hansen solubility parameter of between  $6.0 \text{ (MPa)}^{1/2}$  and  
 $17.0 \text{ (MPa)}^{1/2}$ ; and

10 has a hydrogen bonding Hansen solubility parameter of between  $4.0 \text{ (MPa)}^{1/2}$   
and  $13.0 \text{ (MPa)}^{1/2}$ .

53. The process of claim 52 wherein the aprotic solvent:

has an evaporation rate of lower than 50 (based on n-butyl acetate =  
100); and

15 has a flash point greater than 100 degrees Fahrenheit.

54. The process of claim 1 wherein the pressurized fluid solvent is  
densified carbon dioxide.

20 55. A system for cleaning substrates comprising:

a cleaning vessel adapted to hold contaminated substrates and organic  
solvent;

an organic solvent tank operatively connected to the cleaning vessel;

25 a pump for pumping organic solvent from the organic solvent tank to  
the cleaning vessel;

a drying vessel adapted to hold cleaned substrates and pressurized  
fluid solvent;

a pressurized fluid solvent tank operatively connected to the drying  
vessel; and

30 a pump for pumping pressurized fluid solvent from the pressurized fluid  
solvent tank to the drying vessel.

56. The system of claim 55 wherein the organic solvent comprises a cyclic  
terpene.



57. The system of claim 56 wherein the cyclic terpene:  
is soluble in carbon dioxide between 600 and 1050 pounds per square  
inch and between 5 and 30 degrees Celsius;  
has a specific gravity of greater than approximately 0.800;  
has a dispersion Hansen solubility parameter of between  $13.0 \text{ (MPa)}^{1/2}$   
and  $17.5 \text{ (MPa)}^{1/2}$ ;  
has a polar Hansen solubility parameter of between  $0.5 \text{ (MPa)}^{1/2}$  and  
 $9.0 \text{ (MPa)}^{1/2}$ ; and  
has a hydrogen bonding Hansen solubility parameter of between  $0.0$   
 $\text{ (MPa)}^{1/2}$  and  $10.5 \text{ (MPa)}^{1/2}$ .

58. The system of claim 57 wherein the cyclic terpene further:  
has an evaporation rate of lower than 50 (based on n-butyl acetate =  
100); and  
has a flash point greater than 100 degrees Fahrenheit.

59. The system of claim 58 wherein the cyclic terpene is selected from a  
group including  $\alpha$ -terpene isomers; pine oil;  $\alpha$ -pinene isomers; d-limonene; and  
mixtures thereof.

60. The system of claim 55 wherein the organic solvent comprises a  
halocarbon.

61. The system of claim 60 wherein the halocarbon:  
is soluble in carbon dioxide between 600 and 1050 pounds per square  
inch and between 5 and 30 degrees Celsius;  
has a specific gravity of greater than approximately 1.100;  
has a dispersion Hansen solubility parameter of between  $10.0 \text{ (MPa)}^{1/2}$   
and  $17.0 \text{ (MPa)}^{1/2}$ ;  
has a polar Hansen solubility parameter of between  $0.0 \text{ (MPa)}^{1/2}$  and  
 $7.0 \text{ (MPa)}^{1/2}$ ; and  
has a hydrogen bonding Hansen solubility parameter of between  $0.0$   
 $\text{ (MPa)}^{1/2}$  and  $5.0 \text{ (MPa)}^{1/2}$ .

62. The system of claim 61 wherein the halocarbon further:  
has an evaporation rate of lower than 50 (based on n-butyl acetate =  
100); and  
has a flash point greater than 100 degrees Fahrenheit.

63. The system of claim 62 wherein the halocarbon is selected from a  
group including chlorinated hydrocarbons; fluorinated hydrocarbons; brominated  
hydrocarbons; and mixtures thereof.

64. The system of claim 55 wherein the organic solvent comprises a glycol  
ether.

65. The system of claim 64 wherein the glycol ether:  
is soluble in carbon dioxide between 600 and 1050 pounds per square  
inch and between 5 and 30 degrees Celsius;  
has a specific gravity of greater than approximately 0.800;  
has a dispersion Hansen solubility parameter of between  $13.0 \text{ (MPa)}^{1/2}$   
and  $19.5 \text{ (MPa)}^{1/2}$ ;  
has a polar Hansen solubility parameter of between  $3.0 \text{ (MPa)}^{1/2}$  and  
 $7.5 \text{ (MPa)}^{1/2}$ ; and  
has a hydrogen bonding Hansen solubility parameter of between  $8.0 \text{ (MPa)}^{1/2}$   
and  $17.0 \text{ (MPa)}^{1/2}$ .

66. The system of claim 65 wherein the glycol ether further:  
has an evaporation rate of lower than 50 (based on n-butyl acetate =  
100); and  
has a flash point greater than 100 degrees Fahrenheit.

67. The system of claim 66 wherein the glycol ether is selected from a  
group including monoethylene glycol ether; diethylene glycol ether; triethylene glycol  
ether; monopropylene glycol ether; dipropylene glycol ether; tripropylene glycol  
ether; and mixtures thereof.

68. The system of claim 55 wherein the organic solvent comprises a polyol.

69. The system of claim 68 wherein the polyol:  
is soluble in carbon dioxide between 600 and 1050 pounds per square  
inch and between 5 and 30 degrees Celsius;  
has a specific gravity of greater than approximately 0.920;  
has a dispersion Hansen solubility parameter of between  $14.0 \text{ (MPa)}^{1/2}$   
and  $18.2 \text{ (MPa)}^{1/2}$ ;  
has a polar Hansen solubility parameter of between  $4.5 \text{ (MPa)}^{1/2}$  and  
 $20.5 \text{ (MPa)}^{1/2}$ ; and  
has a hydrogen bonding Hansen solubility parameter of between  $15.0 \text{ (MPa)}^{1/2}$  and  $30.0 \text{ (MPa)}^{1/2}$ .

70. The system of claim 69 wherein the polyol further:  
has an evaporation rate of lower than 50 (based on n-butyl acetate =  
100); and  
has a flash point greater than 100 degrees Fahrenheit.

71. The system of claim 70 wherein the polyol contains two or more  
hydroxyl radicals.

72. The system of claim 55 wherein the organic solvent comprises an  
ether.

73. The system of claim 72 wherein the ether:  
is soluble in carbon dioxide between 600 and 1050 pounds per square  
inch and between 5 and 30 degrees Celsius;  
has a specific gravity of greater than approximately 0.800;  
has a dispersion Hansen solubility parameter of between  $14.5 \text{ (MPa)}^{1/2}$   
and  $20.0 \text{ (MPa)}^{1/2}$ ;  
has a polar Hansen solubility parameter of between  $1.5 \text{ (MPa)}^{1/2}$  and  
 $6.5 \text{ (MPa)}^{1/2}$ ; and  
has a hydrogen bonding Hansen solubility parameter of between  $5.0 \text{ (MPa)}^{1/2}$  and  $10.0 \text{ (MPa)}^{1/2}$ .

74. The system of claim 73 wherein the ether further:  
has an evaporation rate of lower than 50 (based on n-butyl acetate =  
100); and  
has a flash point greater than 100 degrees Fahrenheit.

75. The system of claim 74 wherein the ether contains no free hydroxyl radicals.

76. The system of claim 55 wherein the organic solvent comprises an ester of glycol ethers.

77. The system of claim 76 wherein the ester of glycol ethers:  
is soluble in carbon dioxide between 600 and 1050 pounds per square  
inch and between 5 and 30 degrees Celsius;  
has a specific gravity of greater than approximately 0.800;  
has a dispersion Hansen solubility parameter of between 15.0 (MPa)<sup>1/2</sup>  
and 20.0 (MPa)<sup>1/2</sup>;  
has a polar Hansen solubility parameter of between 3.0 (MPa)<sup>1/2</sup> and  
10.0 (MPa)<sup>1/2</sup>; and  
has a hydrogen bonding Hansen solubility parameter of between 8.0  
(MPa)<sup>1/2</sup> and 16.0 (MPa)<sup>1/2</sup>.

78. The system of claim 77 wherein the ester of glycol ethers further:  
has an evaporation rate of lower than 50 (based on n-butyl acetate =  
100); and  
has a flash point greater than 100 degrees Fahrenheit.

79. The system of claim 55 wherein the organic solvent comprises an ester of monobasic carboxylic acids.

80. The system of claim 79 wherein the ester of monobasic carboxylic acids:

is soluble in carbon dioxide between 600 and 1050 pounds per square inch and between 5 and 30 degrees Celsius;

has a specific gravity of greater than approximately 0.800;

has a dispersion Hansen solubility parameter of between  $13.0 \text{ (MPa)}^{1/2}$  and  $17.0 \text{ (MPa)}^{1/2}$ ;

has a polar Hansen solubility parameter of between  $2.0 \text{ (MPa)}^{1/2}$  and  $7.5 \text{ (MPa)}^{1/2}$ ; and

has a hydrogen bonding Hansen solubility parameter of between  $1.5 \text{ (MPa)}^{1/2}$  and  $6.5 \text{ (MPa)}^{1/2}$ .

81. The system of claim 80 wherein the ester of monobasic carboxylic acids further:

has an evaporation rate of lower than 50 (based on n-butyl acetate = 100); and

has a flash point greater than 100 degrees Fahrenheit.

82. The system of claim 55 wherein the organic solvent comprises a fatty alcohol.

83. The system of claim 82 wherein the fatty alcohol:

is soluble in carbon dioxide between 600 and 1050 pounds per square inch and between 5 and 30 degrees Celsius;

has a specific gravity of greater than approximately 0.800;

has a dispersion Hansen solubility parameter of between  $13.3 \text{ (MPa)}^{1/2}$  and  $18.4 \text{ (MPa)}^{1/2}$ ;

has a polar Hansen solubility parameter of between  $3.1 \text{ (MPa)}^{1/2}$  and  $18.8 \text{ (MPa)}^{1/2}$ ; and

has a hydrogen bonding Hansen solubility parameter of between  $8.4 \text{ (MPa)}^{1/2}$  and  $22.3 \text{ (MPa)}^{1/2}$ .

84. The system of claim 83 wherein the fatty alcohol further:

has an evaporation rate of lower than 50 (based on n-butyl acetate = 100); and

has a flash point greater than 100 degrees Fahrenheit.

85. The system of claim 84 wherein, in the fatty alcohol, the carbon chain adjacent to the hydroxyl group contains at least five carbon atoms.

5 86. The system of claim 55 wherein the organic solvent comprises a short chain alcohol.

87. The system of claim 86 wherein the short chain alcohol:  
is soluble in carbon dioxide between 600 and 1050 pounds per square  
10 inch and between 5 and 30 degrees Celsius;  
has a specific gravity of greater than approximately 0.800;  
has a dispersion Hansen solubility parameter of between  $13.5 \text{ (MPa)}^{1/2}$   
and  $18.0 \text{ (MPa)}^{1/2}$ ;  
has a polar Hansen solubility parameter of between  $3.0 \text{ (MPa)}^{1/2}$  and  
15  $9.0 \text{ (MPa)}^{1/2}$ ; and  
has a hydrogen bonding Hansen solubility parameter of between  $9.0 \text{ (MPa)}^{1/2}$  and  $16.5 \text{ (MPa)}^{1/2}$ .

88. The system of claim 87 wherein the short chain alcohol further:  
20 has an evaporation rate of lower than 50 (based on n-butyl acetate =  
100); and  
has a flash point greater than 100 degrees Fahrenheit.

89. The system of claim 88 wherein, in the short chain alcohol, the carbon  
25 chain adjacent to the hydroxyl group contains no more than four carbon atoms.

90. The system of claim 55 wherein the organic solvent comprises a  
siloxane.

30 91. The system of claim 90 wherein the siloxane:  
is soluble in carbon dioxide between 600 and 1050 pounds per square  
inch and between 5 and 30 degrees Celsius;  
has a specific gravity of greater than approximately 0.900;

45

has a dispersion Hansen solubility parameter of between  $14.0 \text{ (MPa)}^{1/2}$  and  $18.0 \text{ (MPa)}^{1/2}$ ;

has a polar Hansen solubility parameter of between  $0.0 \text{ (MPa)}^{1/2}$  and  $4.5 \text{ (MPa)}^{1/2}$ ; and

has a hydrogen bonding Hansen solubility parameter of between  $0.0 \text{ (MPa)}^{1/2}$  and  $4.5 \text{ (MPa)}^{1/2}$ .

92. The system of claim 91 wherein the siloxane:

has an evaporation rate of lower than 50 (based on n-butyl acetate =

100); and

has a flash point greater than 100 degrees Fahrenheit.

93. The system of claim 55 wherein the organic solvent comprises a hydrofluoroether.

94. The system of claim 93 wherein the hydrofluoroether:

is soluble in carbon dioxide between 600 and 1050 pounds per square inch and between 5 and 30 degrees Celsius;

has a specific gravity of greater than approximately 1.500;

has a dispersion Hansen solubility parameter of between  $12.0 \text{ (MPa)}^{1/2}$  and  $18.0 \text{ (MPa)}^{1/2}$ ;

has a polar Hansen solubility parameter of between  $4.0 \text{ (MPa)}^{1/2}$  and  $10.0 \text{ (MPa)}^{1/2}$ ; and

has a hydrogen bonding Hansen solubility parameter of between  $1.5 \text{ (MPa)}^{1/2}$  and  $9.0 \text{ (MPa)}^{1/2}$ .

95. The system of claim 94 wherein the hydrofluoroether:

has an evaporation rate of lower than 50 (based on n-butyl acetate =

100); and

has a flash point greater than 100 degrees Fahrenheit.

96. The system of claim 55 wherein the organic solvent comprises an aliphatic hydrocarbon.

97. The system of claim 96 wherein the aliphatic hydrocarbon:  
is soluble in carbon dioxide between 600 and 1050 pounds per square  
inch and between 5 and 30 degrees Celsius;  
has a specific gravity of greater than approximately 0.700;  
5 has a dispersion Hansen solubility parameter of between  $14.0 \text{ (MPa)}^{1/2}$   
and  $17.0 \text{ (MPa)}^{1/2}$ ;  
has a polar Hansen solubility parameter of between  $0.0 \text{ (MPa)}^{1/2}$  and  
 $2.0 \text{ (MPa)}^{1/2}$ ; and  
has a hydrogen bonding Hansen solubility parameter of between  $0.0$   
10  $\text{(MPa)}^{1/2}$  and  $2.0 \text{ (MPa)}^{1/2}$ .

98. The system of claim 97 wherein the aliphatic hydrocarbon:  
has an evaporation rate of lower than 50 (based on n-butyl acetate =  
100); and  
15 has a flash point greater than 100 degrees Fahrenheit.

99. The system of claim 55 wherein the organic solvent comprises an ester  
of dibasic carboxylic acids.

20 100. The system of claim 99 wherein the ester of dibasic carboxylic acids:  
is soluble in carbon dioxide between 600 and 1050 pounds per square  
inch and between 5 and 30 degrees Celsius;  
has a specific gravity of greater than approximately 0.900;  
has a dispersion Hansen solubility parameter of between  $13.5 \text{ (MPa)}^{1/2}$   
25 and  $18.0 \text{ (MPa)}^{1/2}$ ;  
has a polar Hansen solubility parameter of between  $4.0 \text{ (MPa)}^{1/2}$  and  
 $6.5 \text{ (MPa)}^{1/2}$ ; and  
has a hydrogen bonding Hansen solubility parameter of between  $4.0$   
30  $\text{(MPa)}^{1/2}$  and  $11.0 \text{ (MPa)}^{1/2}$ .

101. The system of claim 100 wherein the ester of dibasic carboxylic acids:  
has an evaporation rate of lower than 50 (based on n-butyl acetate =  
100); and  
has a flash point greater than 100 degrees Fahrenheit.



102. The system of claim 55 wherein the organic solvent comprises a ketone.

5 103. The system of claim 102 wherein the ketone:  
is soluble in carbon dioxide between 600 and 1050 pounds per square  
inch and between 5 and 30 degrees Celsius;  
has a specific gravity of greater than approximately 0.800;  
has a dispersion Hansen solubility parameter of between  $13.0 \text{ (MPa)}^{1/2}$   
10 and  $19.0 \text{ (MPa)}^{1/2}$ ;  
has a polar Hansen solubility parameter of between  $3.0 \text{ (MPa)}^{1/2}$  and  
 $8.0 \text{ (MPa)}^{1/2}$ ; and  
has a hydrogen bonding Hansen solubility parameter of between  $3.0$   
 $\text{(MPa)}^{1/2}$  and  $11.0 \text{ (MPa)}^{1/2}$ .

15 104. The system of claim 103 wherein the ketone:  
has an evaporation rate of lower than 50 (based on n-butyl acetate =  
100); and  
has a flash point greater than 100 degrees Fahrenheit.

20 105. The system of claim 55 wherein the organic solvent comprises an  
aprotic solvent that contains no dissociable hydrogens.

25 106. The system of claim 105 wherein the aprotic solvent:  
is soluble in carbon dioxide between 600 and 1050 pounds per square  
inch and between 5 and 30 degrees Celsius;  
has a specific gravity of greater than approximately 0.900;  
has a dispersion Hansen solubility parameter of between  $15.0 \text{ (MPa)}^{1/2}$   
and  $21.0 \text{ (MPa)}^{1/2}$ ;  
30 has a polar Hansen solubility parameter of between  $6.0 \text{ (MPa)}^{1/2}$  and  
 $17.0 \text{ (MPa)}^{1/2}$ ; and  
has a hydrogen bonding Hansen solubility parameter of between  $4.0$   
 $\text{(MPa)}^{1/2}$  and  $13.0 \text{ (MPa)}^{1/2}$ .

107. The system of claim 106 wherein the aprotic solvent:  
has an evaporation rate of lower than 50 (based on n-butyl acetate =  
100); and  
has a flash point greater than 100 degrees Fahrenheit.

5

108. The system of claim 55 wherein the pressurized fluid solvent is  
densified carbon dioxide.

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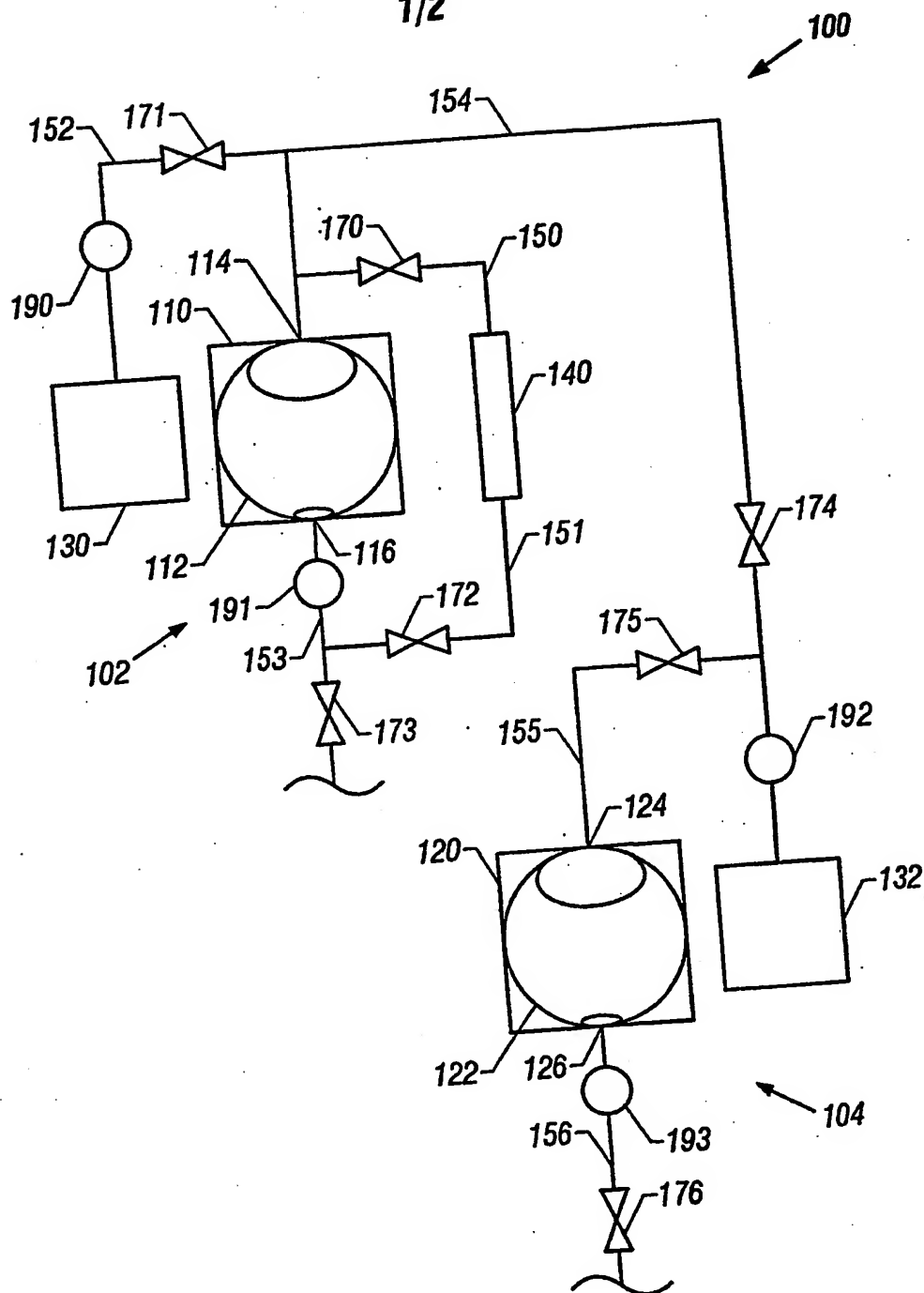
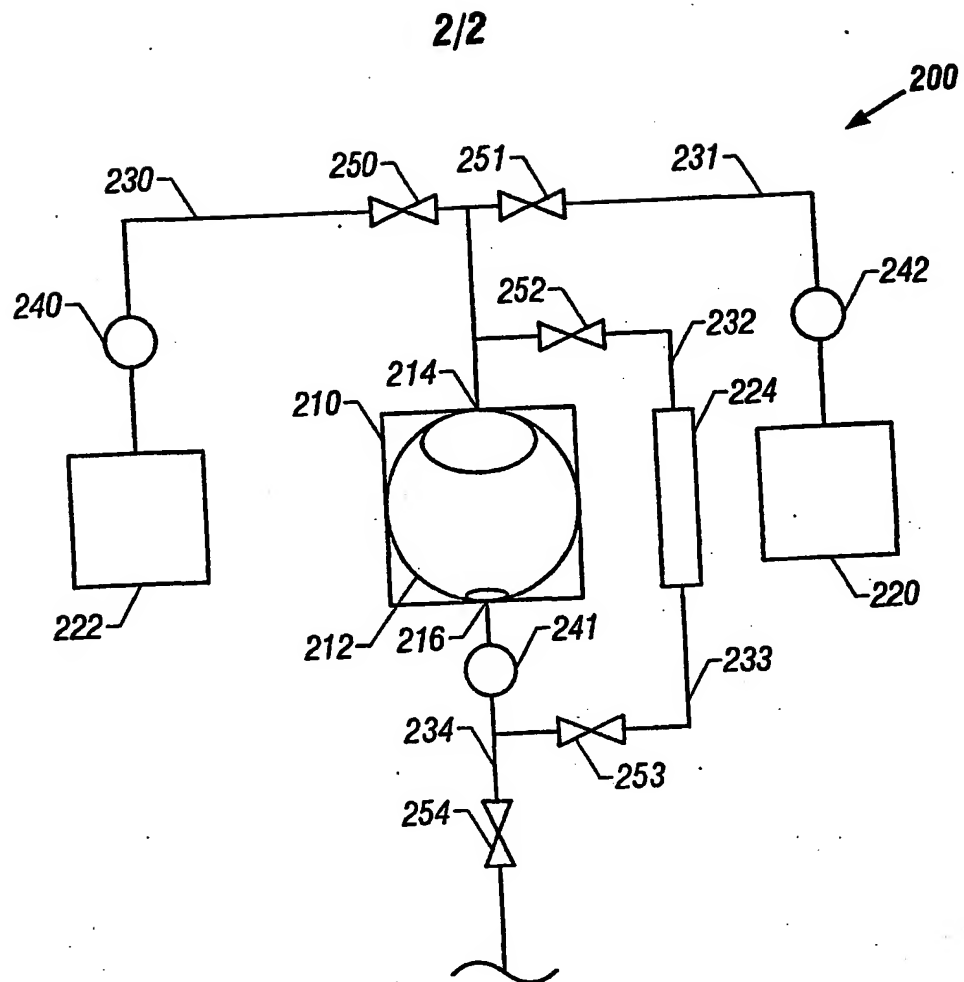


FIG. 1



**FIG. 2**

# INTERNATIONAL SEARCH REPORT

International Application No  
PCT/US 02/12304

A. CLASSIFICATION OF SUBJECT MATTER  
IPC 7 D06L1/02 C11D7/50 B08B7/00

According to International Patent Classification (IPC) or to both national classification and IPC

## B. FIELDS SEARCHED

Minimum documentation searched (classification system followed by classification symbols)  
IPC 7 D06L C11D B08B D06F

Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched

Electronic data base consulted during the international search (name of data base and, where practical, search terms used)

EPO-Internal

## C. DOCUMENTS CONSIDERED TO BE RELEVANT

Category *	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
P, X	WO 01 29305 A (RACETTE ET AL) 26 April 2001 (2001-04-26) the whole document — -/-	1-108

☒ Further documents are listed in the continuation of box C.

☒ Patent family members are listed in annex.

### \* Special categories of cited documents:

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Date of the actual completion of the international search

23 August 2002

Date of mailing of the international search report

02/09/2002

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International Application No  
PCT/US 02/12304

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International Application No

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# INTERNATIONAL SEARCH REPORT

International Application No  
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column 6, line 3 - line 11  
claim 1; figure

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# INTERNATIONAL SEARCH REPORT

International Application No  
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	abstract page 4, line 10 - line 30; claims — -/--	

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International Application No  
PCT/US 02/12304

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103,104,  
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abstract  
column 5, line 33 - line 41  
column 6, line 2 - line 4  
column 6, line 53 - column 7, line 1  
column 8, line 19 - line 24; figure 1  
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# INTERNATIONAL SEARCH REPORT

Int.....ational Application No  
PCT/US 02/12304

C.(Continuation) DOCUMENTS CONSIDERED TO BE RELEVANT		
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